



Lab Manual of **MATHEMATICS**

Class X

Rohit Publications Class 10 Maths Lab Manual

N Colangelo



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 $2 \text{ CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g}) + \text{C}_2\text{H}_2(\text{g}) \rightarrow \text{C}_2\text{H}_6(\text{g}) + \text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g})$
 $\Delta h = 94.5 \text{ kJ}$
 $\text{H}_2\text{O}(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l})$
 $\Delta h = 71.2 \text{ kJ}$
 $\text{C}_2\text{H}_6(\text{g}) + \frac{7}{2} \text{O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 3 \text{ H}_2\text{O}(\text{g})$
 $\Delta h = 283 \text{ kJ}$
 $\text{C}_2\text{H}_6(\text{g})$

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 $\text{C}_2\text{H}_4(\text{g}) + 3 \text{ O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 2 \text{ H}_2\text{O}(\text{l})$
 $\Delta h = 1411 \text{ kJ}$
 $\text{C}_2\text{H}_6(\text{g}) + 7 \text{ O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 3 \text{ H}_2\text{O}(\text{l})$
 $\Delta h = 1560 \text{ kJ}$
 $\text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l})$
 $\Delta h = 285 \text{ kJ}$
 mole

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 $\text{1 NaOH (s)} + \text{Na}^+ \text{aq} + \text{OH}^- \text{aq} \rightarrow \text{H}_2\text{O (l)}$
 $\text{2 NaOH (s)} + \text{H}^+ \text{aq} + \text{Cl}^- \text{aq} \rightarrow \text{H}_2\text{O (l)}$
 $\text{3 NaOH (s)} + \text{H}^+ \text{aq} + \text{Cl}^- \text{aq} \rightarrow \text{H}_2\text{O (l)}$