

Part 3: Regional & Contact Metamorphism: Use the diagram below, your knowledge of earth science and the Earth Science Reference Tables to answer the following questions about metamorphic rocks.



1. Which metamorphic rock formed in the area labeled Continental mountains? \_\_\_\_\_
2. Why do continental mountains cause metamorphic rocks to form? \_\_\_\_\_  
\_\_\_\_\_
3. What type of rock is the granite mountain core made of, metamorphic, sedimentary, or igneous? \_\_\_\_\_  
\_\_\_\_\_
4. What type of metamorphism is occurring near the melted granite core? \_\_\_\_\_
5. Which rock is shown at A in the diagram? Is this a metamorphic, sedimentary, or igneous rock? \_\_\_\_\_  
\_\_\_\_\_
6. Which rock is shown at B in the diagram? Is this a metamorphic, sedimentary, or igneous rock? \_\_\_\_\_  
\_\_\_\_\_
7. Of the metamorphic rocks labelled in the diagram, which will only form from contact metamorphism? \_\_\_\_\_
8. Of the metamorphic rocks labelled in the diagram, which is formed from the low-grade metamorphism of shale? \_\_\_\_\_
9. Of the metamorphic rocks labelled in the diagram, which is formed from the high-grade metamorphism of shale? \_\_\_\_\_

# Rocks Guided And Study Answers Metamorphic Rocks

**Tao Wei**



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web october 3 2017 abstract the techniques used in this lab are quantitative transfer and vacuum filtration with the reaction of 8 001 grams of copper ii sulfate  $\text{CuSO}_4$  and 2 0153 grams of iron powder  $\text{Fe}$  the goal of this experiment was to determine the product of copper ii sulfate with iron

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web procedure 1 find the mass of a clean dry 100 ml beaker and record 2 measure out 8 00 grams of copper ii sulfate onto a piece of filter paper taking the filter paper s mass into account add these to the beaker 3 measure 50 0 ml of water in the cylinder and add it to the beaker 4

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web in this experiment we will use stoichiometric principles to deduce the appropriate equation for the reaction between metallic iron and a solution of copper ii sulfate this reaction produces metallic copper which is seen precipitating as a finely divided red powder

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web in this lab we will use experimental evidence and the principles of stoichiometry to determine which ion is formed in a single replacement reaction between iron and copper ii sulfate pre lab assignment in preparation for the lab activity we will investigate another single replacement reaction between copper and silver nitrate

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web through this experiment we sought to use stoichiometric principles which derive from stoichiometry the study of substance quantities in ratio or proportion in a chemical reaction in this lab experiment the principles help distinguish the equation between iron powder and aqueous copper ii sulfate in the iron copper ii sulfate redox reaction

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web this experiment is focused on a reaction between iron metal and a solution of copper ii sulfate we are able to figure out the chemical equation between the two by using stoichiometric principles the iron displaces the copper ii sulfate ions the

copper ii is then reduced and gains electrons

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web stoichiometry can be defined as the relationship between the relative quantities of substances taking part in a reaction or forming a compound stoichiometric principles can be used to deduce the appropriate chemical equation for a reaction in our experiment we used solid iron in the form of powder was added to a solution of copper ii sulfate

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web under the finalise of the experiment it was found that iron was the limiting reactant and iron ii sulfate was a product of the reaction between 7 0002g away copper ii sulfate and 2 0101g iron metal which produced 2 3037g of copper product with a percent yield of 100 72 introduction

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