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4.14.

$$\left(\frac{101325}{101325}\right)^{\frac{1000}{18}} = 0.5494 \text{ atm}$$

$$\ln \left(\frac{P_{\text{H}_2\text{O}}}{P_{\text{H}_2\text{O}}^*} \right) = -\Delta_{\text{vap}}H_m^\circ / RT = -40.657 \left(\frac{1000}{8.3145} \right) \left(\frac{1}{373} - \frac{1}{298} \right)$$

and that we can combine with the previous:

- 4.15. (a) The water molecule has fewer the external pressure is high.
 (b) The water molecule greater than the external pressure is greater.
 (c) The water molecule greater than the external pressure is approximately 2 atm.

4.16. From the vapor pressure the distribution of the H₂O being P_{H₂O} = P_{H₂O}, the distribution is (101325).

$$\ln \left(\frac{P_{\text{H}_2\text{O}}}{P_{\text{H}_2\text{O}}^*} \right) = -\Delta_{\text{vap}}H_m^\circ / RT = -40.657 \left(\frac{1000}{8.3145} \right) \left(\frac{1}{373} - \frac{1}{298} \right) = -1.943$$

4.17. From the vapor pressure (101325) = 101325 = 101325.

$$\ln \left(\frac{P_{\text{H}_2\text{O}}}{P_{\text{H}_2\text{O}}^*} \right) = -\Delta_{\text{vap}}H_m^\circ / RT = -40.657 \left(\frac{1000}{8.3145} \right) \left(\frac{1}{373} - \frac{1}{298} \right)$$

Therefore

$$\ln \left(\frac{P_{\text{H}_2\text{O}}}{P_{\text{H}_2\text{O}}^*} \right) = -1.943$$

4.18. From the vapor pressure the distribution of the H₂O being P_{H₂O} = P_{H₂O}, the distribution is (101325).

$$\ln \left(\frac{P_{\text{H}_2\text{O}}}{P_{\text{H}_2\text{O}}^*} \right) = -\Delta_{\text{vap}}H_m^\circ / RT = -40.657 \left(\frac{1000}{8.3145} \right) \left(\frac{1}{373} - \frac{1}{298} \right)$$

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Camilla Rothe



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