

Chapter 12: Atoms and the Periodic Table
Subatomic Particles

Three fundamental particles of the atom are the _____, _____, and _____. At the center of each atom lies the atomic _____, which consists of _____ and _____. The atomic number refers to the number of _____ in the nucleus. All atoms of the same element have the same number of _____, hence, the same atomic number.

Isotopes are atoms that have the same number of _____ but a different number of _____. An isotope is identified by its atomic mass number, which is the total number of _____ and _____ in the nucleus. A carbon isotope that has 6 _____ and 6 _____ is identified as carbon-12, where 12 is the atomic mass number. A carbon isotope having 6 _____ and 8 _____, on the other hand, is carbon-14.

1. Complete the following table:

Isotope	Number of...		
	Electrons	Protons	Neutrons
Hydrogen-1	1		
Chlorine-36		17	
Nitrogen-14			7
Potassium-40	19		
Arsenic-75		33	
Gold-197			118

2. Which results in a more valuable product — adding or subtracting protons from gold nuclei?

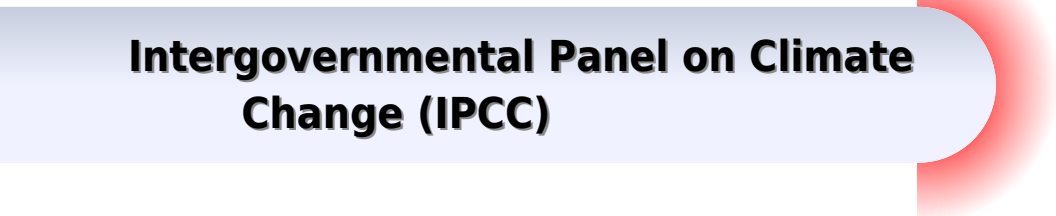
3. Which has more mass, a helium atom or a neon atom?

4. Which has a greater number of atoms, a gram of helium or a gram of neon?



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