

Periodic Trends Worksheet

Directions: Use the table to determine the following properties.

1. Write the following elements by increasing atomic radius: sodium, sulfur, chlorine, oxygen, potassium.

K, Na, O, S, Cl

2. Write the following elements by increasing electronegativity: sodium, oxygen, sulfur, chlorine, potassium.

K, Na, S, O, Cl

3. Which atom has the highest first ionization energy than sodium?

Chlorine (because it has the most electronegativity and is closest to sodium)

4. Which two elements in the same family generally have similar properties?

There have been two atoms that share a same position and a similar properties.

5. Evaluate whether the following properties increase or decrease from left to right across the periodic table:

- a. atomic radius (increases to the right) **decreases**
- b. first ionization energy **increases**
- c. electronegativity **increases**

6. Which period in the periodic table shows atoms with the greatest atomic radius? Which element has the lowest?

Period 7, francium (because it's the furthest from the nucleus)

7. Which period has the lowest average atomic mass compared to the periodic table? Which element has the highest?

Period 1, hydrogen (because it's the closest to the nucleus)

8. Write the atoms in each group that have the highest atomic radius.

1	Li	Be
2	Na	Mg
3	K	Ca
4	Rb	Sr
5	Cs	Ba
6	Fr	Ra

13	B	Al
14	C	Si
15	N	P
16	O	S
17	F	Cl
18	Ar	Kr

9. Circle the atoms in each group that have the most electronegativity.

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