

14

THE BEHAVIOR OF GASES

Chapter Test II

A. Matching

Match a word or phrase in Column I with the correct description in Column II. Write the letter of the correct answer next to the word.

Column I

1. The constant relating unit temperature changes to changes in volume for a fixed amount of gas at constant pressure (the ratio of the change in volume to the change in temperature).
2. The volume of a fixed amount of a gas at constant pressure in the standard state (0°C and 1 atm).
3. The ratio of the mass of a gas to its volume, independent of the actual amount of gas used.
4. An experimental gas constant determined from the ideal gas law at low pressures.
5. A constant of nature that has values of about 8.314 J/K·mol.
6. The a parameter of gas constant equations for determining the gas molecular weight from pressure.
7. An experimental gas constant derived from the ideal gas law under standard conditions (0°C and 1 atm).
8. $\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$
9. The amount of heat enough to raise 1 g of water 1°C.
10. Ratio of volume to mass

Column II

- a. standard gas law
- b. ideal gas constant R
- c. density
- d. compressibility
- e. Boyle's law
- f. molar volume
- g. standard molar volume
- h. volume
- i. compressibility
- j. heat capacity of oxygen

B. Multiple Choice

Choose the correct answer and write its letter on the line.

1. Expanding one volume of a monatomic gas from 1 atm and 27°C to 1 atm and 127°C requires enough energy to
 - a. be determined from Boyle's law
 - b. be calculated by using Charles' law
 - c. be determined by using Boyle's law
 - d. be calculated by using Boyle's law

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CL Gary



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