

**7-3 Review and Reinforcement (continued)**

Write the name of each of the following ionic compounds.

- |                    |                            |       |
|--------------------|----------------------------|-------|
| 17. $ZnO$          | <b>zinc oxide</b>          | _____ |
| 18. $NaNO_3$       | <b>sodium nitrate</b>      | _____ |
| 19. $K_2CO_3$      | <b>potassium carbonate</b> | _____ |
| 20. $Mg(OH)_2$     | <b>magnesium hydroxide</b> | _____ |
| 21. $K_2S$         | <b>potassium sulfide</b>   | _____ |
| 22. $CaF$          | <b>calcium fluoride</b>    | _____ |
| 23. $AgCl$         | <b>silver chloride</b>     | _____ |
| 24. $(NH_4)_3PO_4$ | <b>ammonium phosphate</b>  | _____ |

Write the name of each of the following molecular compounds.

- |              |                                 |       |
|--------------|---------------------------------|-------|
| 25. $SO_2$   | <b>sulfur dioxide</b>           | _____ |
| 26. $HI$     | <b>hydroiodic acid</b>          | _____ |
| 27. $P_2S_5$ | <b>diphosphorus trisulfide</b>  | _____ |
| 28. $NF_3$   | <b>nitrogen trifluoride</b>     | _____ |
| 29. $SF_6$   | <b>sulfur hexafluoride</b>      | _____ |
| 30. $H_2S$   | <b>hydrosulfuric acid</b>       | _____ |
| 31. $N_2O_5$ | <b>dinitrogen pentoxide</b>     | _____ |
| 32. $PCl_5$  | <b>phosphorus pentachloride</b> | _____ |

# Prentice Hall Chemistry Answer Key Ch 22

**Gerald A. Guter, G. S. Hammond**



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Skoog, Donald M. West, F. James Holler, 1992      **The Freeman Library of Laboratory Separates in Chemistry:**

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Lemay, Prentice-Hall Staff, 1996      **Standard Potentials in Aqueous Solution** Allen J. Bard, Roger Parsons, Joseph Jordan, 2017-11-22 The best available collection of thermodynamic data The first of its kind in over thirty years this up to date book presents the current knowledge on Standard Potentials in Aqueous Solution Written by leading international experts and initiated by the IUPAC Commissions on Electrochemistry and Electroanalytical Chemistry this remarkable work begins with a thorough review of basic concepts and methods for determining standard electrode potentials Building upon this solid foundation this convenient source proceeds to discuss the various redox couples for every known element The chapters of this practical time saving guide are organized in order of the groups of elements on the periodic table for easy reference to vital material AND each chapter also contains the fundamental chemistry of elements numerous equations of chemical reactions easy to read tables of thermodynamic data and useful oxidation state diagrams Standard Potentials in Aqueous Solution is an ideal handy reference for analytical and physical chemists electrochemists electroanalytical chemists chemical engineers biochemists inorganic and organic chemists and spectroscopists needing information on reactions and thermodynamic data in inorganic chemistry And it is a valuable supplementary text for undergraduate and graduate level chemistry students

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**Chemistry of +1 Iodine in Alkaline Solution** Yuan-tsan Chia, 1958 [Ionic Soft Matter: Modern Trends in Theory and Applications](#) Douglas Henderson, Myroslav Holovko, Andrij Trokhymchuk, [Solution Thermodynamics and its Application to Aqueous Solutions](#) Yoshikata Koga, 2007-11-12 As the title suggests we introduce a novel differential approach to solution thermodynamics and use it for the study of aqueous solutions We evaluate the quantities of higher order derivative than the normal thermodynamic functions We allow these higher derivative data speak for themselves without resorting to any model system We thus elucidate the molecular processes in solution referred to in this book mixing scheme to the depth equal to if not deeper than that gained by spectroscopic and other methods We show that there are three composition regions in aqueous solutions of non electrolytes each of which has a qualitatively distinct mixing scheme The boundary between the adjacent

regions is associated with an anomaly in the third derivatives of  $G$ . The loci of the anomalies in the temperature composition field form the line sometimes referred to as Koga line. We then take advantage of the anomaly of a third derivative quantity of 1-propanol in the ternary aqueous solution 1-propanol sample species  $H_2O$ . We use its induced change as a probe of the effect of a sample species on  $H_2O$ . In this way we clarified what a hydrophobe or a hydrophile and in turn an amphiphile does to  $H_2O$ . We also apply the same methodology to ions that have been ranked by the Hofmeister series. We show that the kosmotropes salting out or stabilizing agents are either hydrophobes or hydration centres and that chaotropes salting in or destabilizing agents are hydrophiles. A new differential approach to solution thermodynamics. A particularly clear elucidation of the mixing schemes in aqueous solutions. A clear understanding on the effects of hydrophobes, hydrophiles and amphiphiles to  $H_2O$ . A clear understanding on the effects of ions on  $H_2O$  in relation to the Hofmeister effect. A new differential approach to studies in multi-component aqueous solutions.

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Physical Chemistry Peter Atkins, Julio de Paula, 2006-03-10 Change 21

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