

Chapter 7 – Chemical Quantities and Reactions

7.1 THE MOLE

Collection Term // States a specific number of items.

- 1 dozen donuts = 12 donuts
- 1 ream of paper = 500 sheets
- 1 case = 24 cans

Mole // A collection that contains the same number of particles as there are carbon atoms in 12.0 g of carbon: 6.02×10^{23} atoms of an element.

Avogadro's Number // 6.02×10^{23} atoms of an element.

- 1 mole = 6.02×10^{23} atoms/molecules & this is an equality
- Used to convert moles of a substance to particles
- SIG FIGS count

Subscripts and Moles

- The subscripts in a formula show...
 - The relationship of atoms in the formula
 - The moles of each element in 1 mole of compound
- The subscripts are used to write conversion factors for moles of each element in 1 mole of a compound
 - 9 moles C/1 mole $C_9H_8O_4$
 - 8 moles H/1 mole $C_9H_8O_4$
 - 4 moles O/1 mole $C_9H_8O_4$

7.2-3 MOLAR MASS & CALCULATIONS

Molar Mass // The mass of 1 mole of an element or compound (atomic mass expressed in g).

- For a compound...
 - The sum of the molar masses of the elements in the formula
 - $CaCl_2 = 40.1g/mol (Ca) + (2 \times 35.5g/mol) (Cl)$

Molar Mass Conversion Factors

- Are fractions written from the molar mass

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