

REVIEW: Chapter 8 - Covalent Bonding

- Typically, how many bonds & lone pairs will each have when covalently bonding?
 - Boron – 3 bonds; 0 lone pairs
 - Silicone – 4 bonds; 0 lone pairs
 - Neon – 0 bonds; 4 lone pairs
 - Iodine – 1 bond; 3 lone pairs
 - Nitrogen – 3 bonds; 1 lone pair
 - Oxygen – 2 bonds; 2 lone pairs
- Which elements do not follow the octet rule? (Hint – there are four we went over)
 - ___H___ & ___B___ never follow the octet rule.
 - ___P___ & ___S___ sometimes don't follow the octet rule.
- What type of bond will each combination of atoms form?
 - F & N – 2 nonmetals, therefore covalent
 - Al & O – 1 metal/1 nonm., therefore ionic
 - Na & S – 1 metal/1 nonm., therefore ionic
 - B & Cl – 2 nonmetals, therefore covalent
 - Ca & I – 1 metal/1 nonm., therefore ionic
 - Fe & Cu – 2 metals, therefore metallic
- What is the difference between:
 - Molecules & Compounds?
Compounds can be made using all types of bonds (ionic, covalent or metallic), but molecules can only be made using covalent bonds
 - Monatomic elements & Diatomic molecules ?
Monatomic elements exist as single atoms; diatomic molecules are made up of two of the same atom
 - Structural formulas, Empirical formulas, & Molecular formulas?
All three are for covalent bonds only! Molecular show the actual number of atoms in the molecule; Empirical give the lowest whole number ratio; Structural formulas are not really formulas, they are drawing of the molecule.
 - Polar bonds & Non-polar bonds?
Both describe covalent bonds! In a polar bond the electrons are NOT shared equally; in a non-polar bond electrons are shared equally
- Classify each of the following as an atom or a molecule:
 - Be, atom
 - CO, molec.
 - N₂, molec.
 - H₂O, molec.
 - Ne, atom

Prentice Hall Chemistry Chapter 8 Covalent Bonding Answers

James C. Hill, Bruce Edward Bursten



Prentice Hall Chemistry Chapter 8 Covalent Bonding Answers:

Adopting the Tune of Phrase: An Psychological Symphony within **Prentice Hall Chemistry Chapter 8 Covalent Bonding Answers**

In some sort of consumed by screens and the ceaseless chatter of quick interaction, the melodic beauty and mental symphony produced by the prepared word usually disappear in to the back ground, eclipsed by the constant noise and interruptions that permeate our lives. However, located within the pages of **Prentice Hall Chemistry Chapter 8 Covalent Bonding Answers** an enchanting literary treasure full of fresh feelings, lies an immersive symphony waiting to be embraced. Constructed by an elegant musician of language, this captivating masterpiece conducts viewers on a mental trip, skillfully unraveling the concealed songs and profound affect resonating within each cautiously crafted phrase. Within the depths with this emotional analysis, we will investigate the book is main harmonies, analyze its enthralling publishing design, and submit ourselves to the profound resonance that echoes in the depths of readers souls.

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