

14

THE BEHAVIOR OF GASES

Chapter Test II

A. Matching

Match a word or phrase in Column I with the correct description in Column II. Write the letter of the correct description next to the item.

Column I

1. The constant relating unit temperature changes to changes in volume for a fixed amount of gas at constant pressure (the ratio of the change in volume to the change in temperature).
2. The volume of a fixed amount of a gas at constant pressure in the standard temperature and pressure (STP) volume.
3. The ratio of volume of a gas to pressure, proportional to the square root of molecular mass.
4. Temperature at which a fixed amount of gas exerts zero pressure at that volume.
5. A measure of how much the volume of a fixed amount of gas changes.
6. The a parameter of gas pressure dependence for compressing gas isothermally with pressure.
7. Temperature at which a fixed amount of gas exerts zero pressure at that volume.
8. $\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$
9. A measure of how much a fixed amount of gas changes.
10. Ratio of volume to mass.

Column II

- a. standard pressure
- b. ideal gas constant R
- c. volume
- d. compressibility
- e. Boyle's law
- f. Boyle's constant
- g. constant-volume molar heat capacity
- h. volume
- i. compressibility
- j. compressibility coefficient

B. Multiple Choice

Choose the correct answer and write its letter on the line.

1. Expanding one volume of a monatomic gas from 1 atm and 27°C to 1 atm and 127°C requires heating. The expansion is done against a constant pressure of:
a. the atmospheric pressure
b. the pressure of the gas itself
c. the pressure of the wall of the container
d. the pressure of the gas itself

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**Arthur Newell Strahler, Alan H.
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1. Understanding the eBook Prentice Hall Chemistry Chapter 13 And 14 Test Answers
 - The Rise of Digital Reading Prentice Hall Chemistry Chapter 13 And 14 Test Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Prentice Hall Chemistry Chapter 13 And 14 Test Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Prentice Hall Chemistry Chapter 13 And 14 Test Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Prentice Hall Chemistry Chapter 13 And 14 Test Answers
 - Personalized Recommendations
 - Prentice Hall Chemistry Chapter 13 And 14 Test Answers User Reviews and Ratings
 - Prentice Hall Chemistry Chapter 13 And 14 Test Answers and Bestseller Lists
5. Accessing Prentice Hall Chemistry Chapter 13 And 14 Test Answers Free and Paid eBooks
 - Prentice Hall Chemistry Chapter 13 And 14 Test Answers Public Domain eBooks
 - Prentice Hall Chemistry Chapter 13 And 14 Test Answers eBook Subscription Services
 - Prentice Hall Chemistry Chapter 13 And 14 Test Answers Budget-Friendly Options
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- ePub, PDF, MOBI, and More
 - Prentice Hall Chemistry Chapter 13 And 14 Test Answers Compatibility with Devices
 - Prentice Hall Chemistry Chapter 13 And 14 Test Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Prentice Hall Chemistry Chapter 13 And 14 Test Answers
 - Highlighting and Note-Taking Prentice Hall Chemistry Chapter 13 And 14 Test Answers
 - Interactive Elements Prentice Hall Chemistry Chapter 13 And 14 Test Answers
 8. Staying Engaged with Prentice Hall Chemistry Chapter 13 And 14 Test Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Prentice Hall Chemistry Chapter 13 And 14 Test Answers
 9. Balancing eBooks and Physical Books Prentice Hall Chemistry Chapter 13 And 14 Test Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Prentice Hall Chemistry Chapter 13 And 14 Test Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Prentice Hall Chemistry Chapter 13 And 14 Test Answers
 - Setting Reading Goals Prentice Hall Chemistry Chapter 13 And 14 Test Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Prentice Hall Chemistry Chapter 13 And 14 Test Answers
 - Fact-Checking eBook Content of Prentice Hall Chemistry Chapter 13 And 14 Test Answers
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
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