

Chemistry--Unit 10: Chemical Periodicity
Practice Problems

I. Classification of the elements

- 1) Use the periodic table to write the electron configuration for silicon.

$1s^2 2s^2 2p^6 3s^2 3p^2$ or abbreviated form*: $[\text{Ne}] 3s^2 3p^2$ *may use this form now!!!

- 2) Use the periodic table to write the electron configuration for iodine.

$1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^5$ or $[\text{Kr}] 5s^2 4d^{10} 5p^5$

- 3) Which family of elements is characterized by an $s^2 p^3$ configuration?

Nitrogen Family, or Group 15

- 4) Name the element that matches the following description:

- one that has 5 outer electrons on the third period of the periodic table **P, phosphorus**
- one with a $4s^2 4p^5$ electron configuration **Br, bromine**
- the 4th period Group 16 element **Se, selenium**

- 5) Identify the elements that have the following outer electron configurations?

- $2s^2 2p^4$ **O, oxygen**
- $4s^2$ **Ca, Calcium**
- $4s^2 3d^{10}$ **Zn, zinc**

- 6) What is the common characteristic of the electron configurations of the elements Ne and Ar? In which Group would you find them?

full outermost energy level, Group 18, the noble or inert gases

- 7) Why would you expect lithium (Li) and sulfur (S) to have different chemical and physical properties?

they are in different groups on the periodic table (Li group 1, S group 16)

- 8) What characterizes silver (Ag) and iron (Fe) as transition metals?

Their electron configurations both end in the d sublevel (last electrons added are to a d sublevel).

II. Periodic Trends

- 9) Explain why a magnesium atom is smaller than atoms of both sodium and calcium.

Na has less +/- attraction; Ca has another energy level

- 10) Predict the size of the astatine (At) atom compared to that of tellurium (Te).

At would be larger than Te due to the extra energy level.

Prentice Hall Chemistry 3 Practice Problems Answers

Irwin H. Segel



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as Frameworks for Student Learning Tim Riordan,James Roth,2023-07-03 What should students be able to do and how should they be able to think as a result of study in a discipline What does learning in the disciplines look like at different developmental levels How does one go about designing such learning and assessment in the disciplines What institutional structures and processes can assist faculty to engage and teach their disciplines as frameworks for student learning Creating ways to make a discipline come alive for those who are not experts even for students who may not take more than one or two

courses in the disciplines they study requires rigorous thought about what really matters in a field and how to engage students in the practice of it Faculty from Alverno College representing a range of liberal arts disciplines chemistry economics history literature mathematics and philosophy here reflect on what it has meant for them to approach their disciplines as frameworks for student learning They present the intellectual biographies of their explorations the insights they have gained and examples of the practices they have adopted The authors all demonstrate how the ways of thinking they have identified as significant for their students in their respective disciplines have affected the way they design learning experiences and assessments They show how they have shaped their teaching around the ways of thinking they want their students to develop within and across their disciplines and what that means in terms of designing assessments that require students to demonstrate their thinking and understanding through application and use This book will appeal to faculty interested in going beyond mere techniques to a more substantive analysis of how their view of their respective disciplines might change when seen through the lens of student learning It will also serve the needs of graduate students trainers of Tas and anyone engaged in faculty development or interested in the scholarship of teaching

Physical and Chemical Equilibrium for Chemical Engineers Noel de Nevers, 2012-03-20 This book concentrates on the topic of physical and chemical equilibrium Using the simplest mathematics along with numerous numerical examples it accurately and rigorously covers physical and chemical equilibrium in depth and detail It continues to cover the topics found in the first edition however numerous updates have been made including Changes in naming and notation the first edition used the traditional names for the Gibbs Free Energy and for Partial Molal Properties this edition uses the more popular Gibbs Energy and Partial Molar Properties changes in symbols the first edition used the Lewis Randal fugacity rule and the popular symbol for the same quantity this edition only uses the popular notation and new problems have been added to the text Finally the second edition includes an appendix about the Bridgman table and its use

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British Paperbacks in Print, 1985

Learning with Understanding in the Chemistry Classroom Iztok Devetak, Saša Aleksij Glazar, 2014-01-14 This volume offers a critical examination of a variety of conceptual approaches to teaching and learning chemistry in the school classroom Presenting up to date research and theory and featuring

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