

10-3 Practice Problems

- Find the percentage composition of a compound that contains 1.94-g of carbon, 0.48-g of hydrogen, and 2.58-g of sulfur in a 5.00-g sample of the compound.
- A sample of an unknown compound with a mass of 0.847-g has the following composition: 50.51 percent fluorine and 49.49 percent iron. When this compound is decomposed into its elements, what mass of each element would be recovered?
- Find the percentage composition of a compound that contains 2.63-g of carbon, 0.370-g of hydrogen, and 0.580-g of oxygen in a 3.58-g sample of the compound.
- A sample of an unknown compound with a mass of 2.876-g has the following composition: 66.07 percent carbon, 6.71 percent hydrogen, 4.06 percent nitrogen, and 23.16 percent oxygen. What is the mass of each element in this compound?
- Find the percentage composition of a compound that contains 2.7368-g of chlorine, 0.4116-g of oxygen, and 0.7971-g of phosphorus in a 3.9460-g sample of the compound.
- Find the percentage composition of a compound that contains 1.51-g of chromium, 1.13-g of potassium, and 1.82-g of oxygen in a 4.26-g sample of the compound.
- A sample of a compound that has a mass of 0.632-g is analyzed. The sample is found to be made up of oxygen and fluorine only. Given that the sample contains 0.128-g of oxygen, calculate the percentage composition of the compound.
- What is the percentage composition of a carbon-oxygen compound, given that a 95.2-g sample of the compound contains 40.8-g of carbon and 54.4-g of oxygen?
- What is the percentage composition of a sulfur-chlorine compound, given that a 30.9-g sample of the compound contains 9.63-g of sulfur and 21.3-g of chlorine?
- Determine the empirical formula of a compound containing 2.644-g of gold and 0.476-g of chlorine.
- Determine the empirical formula of a compound containing 0.928-g of gallium and 0.612-g of phosphorus.
- Determine the empirical formula of a compound containing 1.723-g of carbon, 0.289-g of hydrogen, and 0.459-g of oxygen.

Prentice Hall 10 3 Practice Problems Key

M Walker



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