

AP Chemistry Chapter 20 Electrochemistry Practice Free Response #1



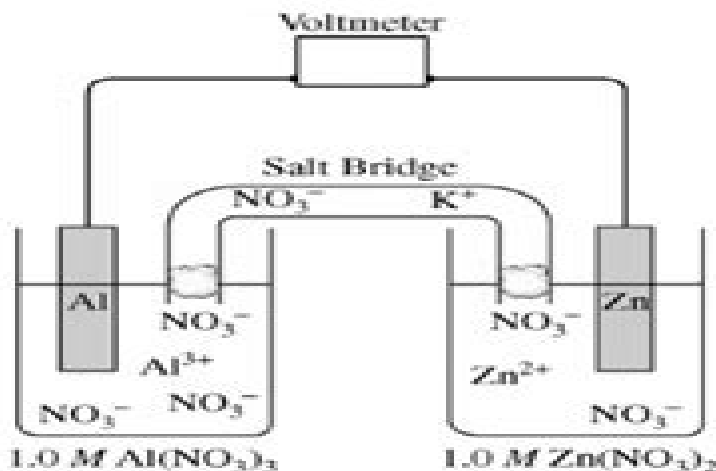
Respond to the following statements and questions that relate to the species and the reaction represented above.

(a) Write the complete electron configuration (e.g., $1s^2 2s^2 \dots$) for Zn^{2+} .

(b) Which species, Zn or Zn^{2+} , has the greater ionization energy? Justify your answer.

(c) Identify the species that is oxidized in the reaction.

The diagram shows a galvanic cell based on the reaction. Assume that the temperature is 25°C .



(d) The diagram includes a salt bridge that is filled with a saturated solution of KNO_3 . Describe what happens in the salt bridge as the cell operates.

(e) Determine the value of the standard voltage, E° , for the cell.

Practice Hall Chemistry Chapter 20

Clifford Lane



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