

## Covalent Bonding Worksheet

Covalent bonding occurs when two or more NON\_METALS share electrons, attempting to attain a stable octet (8 outer electrons) in their outer shell for at least part of the time. Draw a Lewis dot diagram for each element listed. Circle the unpaired electrons that will be shared between the elements.

|                                                                                                                                                                                |
|--------------------------------------------------------------------------------------------------------------------------------------------------------------------------------|
| 1.) H <sub>2</sub> hydrogen is diatomic:<br>$\text{H} + \text{H} \longrightarrow \text{H} \cdot \cdot \text{H} \longrightarrow \text{H}-\text{H} \quad \text{Single Bond}$     |
| 2.) F <sub>2</sub> fluorine is diatomic:<br>$\text{F} + \text{F} \longrightarrow \cdot \ddot{\text{F}} \cdot \cdot \ddot{\text{F}} \cdot \longrightarrow$                      |
| 3.) O <sub>2</sub> oxygen is diatomic:<br>$\text{O} + \text{O} \longrightarrow \quad \quad \quad \longrightarrow \text{O}=\text{O} \quad \text{Double Bond}$                   |
| 4.) N <sub>2</sub> nitrogen is diatomic. Is this a triple bond?<br>$\text{N} + \text{N}$                                                                                       |
| 5.) BF <sub>3</sub> you need 3 fluorine atoms here:<br>$\text{B} + \text{F}$                                                                                                   |
| 6.) Ammonia NH <sub>3</sub> hint: how many hydrogen atoms are needed?<br>$\text{N} + \text{H}$                                                                                 |
| 7.) Carbon dioxide CO <sub>2</sub><br>$\text{O} + \text{C} + \text{O}$                                                                                                         |
| 8.) Methane CH <sub>4</sub> careful here 4 hydrogen atoms needed<br>$\text{C} + \text{H}$                                                                                      |
| 9.) Dihydrogen monoxide: the most dangerous substance on the planet. It has killed more people than any other substance known to mankind!!<br>$\text{H} + \text{O} + \text{H}$ |
| 10.) SO <sub>2</sub> hint: one pair of electrons from sulfur must be slit up for this one to work.<br>$\text{O} + \text{S} + \text{O}$                                         |

# Practice Problems Chemical Bonding

**Arshad Iqbal**



## **Practice Problems Chemical Bonding:**

The book delves into Practice Problems Chemical Bonding. Practice Problems Chemical Bonding is an essential topic that needs to be grasped by everyone, ranging from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Practice Problems Chemical Bonding, encompassing both the fundamentals and more intricate discussions.

1. The book is structured into several chapters, namely:

- Chapter 1: Introduction to Practice Problems Chemical Bonding
- Chapter 2: Essential Elements of Practice Problems Chemical Bonding
- Chapter 3: Practice Problems Chemical Bonding in Everyday Life
- Chapter 4: Practice Problems Chemical Bonding in Specific Contexts
- Chapter 5: Conclusion

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6. In chapter 5, the author will draw a conclusion about Practice Problems Chemical Bonding. This chapter will summarize the key points that have been discussed throughout the book.

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