

13

STATES OF MATTER

SECTION 13.1 THE NATURE OF GASES (pages 385–389)

This section introduces the kinetic theory and describes how it applies to gases. It defines gas pressure and explains how temperature is related to the kinetic energy of the particles of a substance.

► Kinetic Theory and a Model for Gases (pages 385–386)

1. The energy an object has because of its motion is called kinetic energy.
2. Circle the letter of each sentence that is true about the assumptions of the kinetic theory concerning gases.
 - a. A gas is composed of particles with insignificant volume that are relatively far apart from each other.
 - b. Strong attractive forces exist between particles of a gas.
 - c. Gases tend to collect near the bottom of a container.
 - d. The paths of uninterrupted travel of particles in a gas are relatively short because the particles are constantly colliding with each other or with other objects.
3. Is the following statement true or false? According to the kinetic theory, collisions between particles in a gas are perfectly elastic because kinetic energy is transferred without loss from one particle to another, and the total kinetic energy remains constant. true

► Gas Pressure (pages 386–387)

4. Gas pressure results from the force exerted by a gas per unit surface area of an object.
5. Simultaneous collisions of billions of particles in a gas with an object result in gas pressure.
6. What force holds the particles of air in Earth's atmosphere? gravity
7. What kind of pressure is measured with a barometer? atmospheric pressure

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Murray Hiebert, Bruce Klatt



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