

Assessment #3 Review

Name: _____

You must show all work for credit.

Due Date: _____ (No late packets will be accepted)

I. Balance Equations*Balance the equations*

- $\text{Mg} + \underline{2} \text{H}_2\text{O} \rightarrow \text{Mg}(\text{OH})_2 + \text{H}_2$
- $\text{Mg} + \underline{2} \text{NaBr} \rightarrow \text{MgBr}_2 + \underline{2} \text{Na}$
- $\text{C}_3\text{H}_8 + \underline{5} \text{O}_2 \rightarrow \underline{3} \text{CO}_2 + \underline{4} \text{H}_2\text{O}$
- $\underline{4} \text{P} + \underline{5} \text{O}_2 \rightarrow \underline{2} \text{P}_2\text{O}_5$
- $\text{Fe}_2(\text{SO}_4)_3 + \underline{6} \text{KOH} \rightarrow \underline{3} \text{K}_2\text{SO}_4 + \underline{2} \text{Fe}(\text{OH})_3$
- $\text{Fe}_2\text{O}_3 + \underline{3} \text{H}_2 \rightarrow \text{Fe} + \underline{3} \text{H}_2\text{O}$

7. How many moles of hydrogen gas react with 1 mole of Fe_2O_3 in reaction #6? 3

Write the balanced equation for problems # 8 and 9:

8. Hydrogen gas and nitrogen monoxide combine to form water and nitrogen gas



9. When sodium phosphate and calcium chloride are mixed they form calcium phosphate and sodium chloride



10. Using the following balanced equation, prove (or explain) the Law of Conservation of Mass.



2 N on left, 2 N on right, 6 H on left, 6 H on right, 2 on left, 2 on right

11. Using the balanced equation shown in problem #10, explain what the coefficients represent.

moles

12. Using the balanced equation shown in problem #10, list the products of the reaction.

NH_3

13. Using the balanced equation shown in problem #10, list the reactants of the reaction.

N_2 , H_2

II. Types of Reactions*Directions: Identify the type of reaction.*

S = synthesis

SD = single displacement

C = combustion

D = decomposition

DD = double displacement



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Table of Contents Prentice Hall Chemistry Final Exam Review Answers

1. Understanding the eBook Prentice Hall Chemistry Final Exam Review Answers
 - The Rise of Digital Reading Prentice Hall Chemistry Final Exam Review Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Prentice Hall Chemistry Final Exam Review Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Prentice Hall Chemistry Final Exam Review Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Prentice Hall Chemistry Final Exam Review Answers
 - Personalized Recommendations
 - Prentice Hall Chemistry Final Exam Review Answers User Reviews and Ratings
 - Prentice Hall Chemistry Final Exam Review Answers and Bestseller Lists

5. Accessing Prentice Hall Chemistry Final Exam Review Answers Free and Paid eBooks
 - Prentice Hall Chemistry Final Exam Review Answers Public Domain eBooks
 - Prentice Hall Chemistry Final Exam Review Answers eBook Subscription Services
 - Prentice Hall Chemistry Final Exam Review Answers Budget-Friendly Options
6. Navigating Prentice Hall Chemistry Final Exam Review Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Prentice Hall Chemistry Final Exam Review Answers Compatibility with Devices
 - Prentice Hall Chemistry Final Exam Review Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Prentice Hall Chemistry Final Exam Review Answers
 - Highlighting and Note-Taking Prentice Hall Chemistry Final Exam Review Answers
 - Interactive Elements Prentice Hall Chemistry Final Exam Review Answers
8. Staying Engaged with Prentice Hall Chemistry Final Exam Review Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Prentice Hall Chemistry Final Exam Review Answers
9. Balancing eBooks and Physical Books Prentice Hall Chemistry Final Exam Review Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Prentice Hall Chemistry Final Exam Review Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Prentice Hall Chemistry Final Exam Review Answers
 - Setting Reading Goals Prentice Hall Chemistry Final Exam Review Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Prentice Hall Chemistry Final Exam Review Answers
 - Fact-Checking eBook Content of Prentice Hall Chemistry Final Exam Review Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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