

Chapter 7 Chemical Reactions

Section 7.2 Types of Reactions

(pages 199–205)

This section discusses how chemical reactions are classified into different types.

Reading Strategy (page 199)

Previewing. Skim the section and begin a concept map like the one below that identifies types of reactions with a general form.

As you read, add the general forms of each type of reaction. For more information on this Reading Strategy, see the **Reading and Study Skills** in the **Skills and Reference Handbook** at the end of your textbook.



Classifying Reactions (pages 199–204)

1. Name five general types of chemical reactions. Synthesis, decomposition, single replacement, double replacement, and combustion.
2. Circle the letter of each equation that represents a synthesis reaction.
 - a. $2\text{Na} + \text{Cl}_2 \longrightarrow 2\text{NaCl}$
 - b. $2\text{NaCl} \longrightarrow 2\text{Na} + \text{Cl}_2$
 - c. $2\text{H}_2\text{O} \longrightarrow 2\text{H}_2 + \text{O}_2$
 - d. $2\text{H}_2 + \text{O}_2 \longrightarrow 2\text{H}_2\text{O}$
3. Is the following sentence true or false? A decomposition reaction is the opposite of a synthesis reaction. true
4. Write the equation for the decomposition of calcium carbonate into calcium oxide and carbon dioxide. $\text{CaCO}_3 \longrightarrow \text{CaO} + \text{CO}_2$
5. Circle the letter of the correct answer. Copper reacts with silver nitrate in a single-replacement reaction. What are the products of this reaction?
 - a. copper(II) nitride and silver oxide
 - b. copper(II) nitrate and silver
 - c. copper(II) oxide and silver nitrate
 - d. copper, nitrogen, and silver oxide

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