2. Metallic magnesium reacts with steam to produce magnesium hydroxide and hydrogen gas.

a. Write a balanced equation for this reaction. (Note: It is very important that you do this step correctly!)

$$Mg_{(s)} + 2H_2O_{(g)} \longrightarrow Mg(OH)_2 + H_2{}_{(g)}$$

b. What type of reaction is this? single replacement

c. If 16.2 g of Mg are heated with 12.0 g of H_2O, which is the limiting reactant?

$$16.2\text{ g Mg} \times \frac{1\text{ mol Mg}}{24.30\text{ g Mg}} \times \frac{2\text{ mol } H_2O}{1\text{ mol Mg}} \times \frac{18.02\text{ g } H_2O}{1\text{ mol } H_2O} = 24.0\text{ g } H_2O$$

There is less H_2O (12.0 g) in our initial amount so

$\boxed{H_2O \text{ is the L.R}}$

d. How many moles of the excess reactant remain after the reaction is complete?

$$16.2\text{ g Mg} \times \frac{1\text{ mol Mg}}{24.30\text{ g Mg}} \times \frac{2\text{ mol } H_2O}{1\text{ mol Mg}} = 1.3\text{ mol } H_2O$$

$$12.0\text{ g } H_2O \times \frac{1\text{ mol } H_2O}{18.02\text{ g } H_2O} = 0.666\text{ mol } H_2O$$

$$1.3\text{ mol } H_2O - 0.666\text{ mol } H_2O = 0.634\text{ mol } H_2O \times \frac{1\text{ mol Mg}}{2\text{ mol } H_2O} = \boxed{0.317\text{ mol Mg}}$$

e. How many grams of each product are formed?

$$12.0\text{ g } H_2O \times \frac{1\text{ mol } H_2O}{18.02\text{ g } H_2O} \times \frac{1\text{ mol } Mg(OH)_2}{2\text{ mol } H_2O} \times \frac{58.32\text{ g } Mg(OH)_2}{1\text{ mol } Mg(OH)_2} = 19.4\text{ g } Mg(OH)_2$$

$$12.0\text{ g } H_2O \times \frac{1\text{ mol } H_2O}{18.02\text{ g } H_2O} \times \frac{1\text{ mol } H_2}{2\text{ mol } H_2O} \times \frac{2.02\text{ g } H_2}{1\text{ mol } H_2} = 6.28\text{ g } H_2$$

e. You perform this reaction, and obtain 16.0 g $Mg(OH)_2$. What is your percent yield?

$$\%\text{ yield} = \frac{\text{actual}}{\text{theoretical}} \times 100 \qquad \frac{16.0\text{ g } Mg(OH)_2}{19.4\text{ g } Mg(OH)_2} \times 100 = 82.5\%$$

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