

Stoichiometry Reflect #2

Learning Target... I can use stoichiometry to calculate amounts

1a. Balance this equation



1b. Determine how many moles of ammonium oxide was used if the student weighed out 55.2 g of it.

$$55.2 \text{ g (NH}_4\text{)}_2\text{O} \left(\frac{1 \text{ mol A}}{52.10 \text{ g}} \right) = \boxed{1.06 \text{ moles (NH}_4\text{)}_2\text{O}}$$

1c. Determine the amount of moles of ammonium nitride produced in this reaction.

$$1.06 \text{ moles A} \left(\frac{2 \text{ mol B}}{3 \text{ mol A}} \right) = \boxed{0.706 \text{ moles (NH}_4\text{)}_3\text{N}}$$

1d. Determine how many grams of ammonium nitride produced in this reaction.

$$0.706 \text{ moles} \left(\frac{108.11 \text{ g}}{1 \text{ mol B}} \right) = \boxed{48.1 \text{ g (NH}_4\text{)}_3\text{N}}$$

2. Using the following chemical reaction:

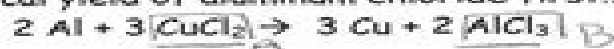


and, knowing that 61.6 g of copper was produced in this experiment, how many molecules of mercury nitrate was also produced?

$$61.6 \text{ g Cu} \left(\frac{1 \text{ mol A}}{63.55 \text{ g}} \right) \left(\frac{2 \text{ mol B}}{1 \text{ mol A}} \right) \left(\frac{6.02 \times 10^{23} \text{ mc}}{1 \text{ mol B}} \right) = \boxed{1.17 \times 10^{24} \text{ mc HgNO}_3}$$

3. Challenge...

What is the percent yield of the following reaction if 12.8 g of copper (II) chloride was used in the lab, and 7.75 g of aluminum chloride was measured at the end of the experiment? (hint: you must calculate the theoretical yield of aluminum chloride first!!)



$$12.8 \text{ g A} \left(\frac{1 \text{ mol A}}{134.45 \text{ g}} \right) \left(\frac{2 \text{ mol B}}{3 \text{ mol A}} \right) \left(\frac{133.33 \text{ g}}{1 \text{ mol B}} \right) = 8.46 \text{ g AlCl}_3$$

theoretical

$$\% \text{ yield} = \frac{7.75 \text{ g}}{8.46 \text{ g}} = \boxed{91.4\%}$$

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Heather Hattori, Richard H. Langley



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CONTENTS 1 Writing the Formulae of Ionic Compounds 2 Formula Masses 3 Mass and Percent Compositions 4 The Mole and Mass 5 The Mole and Number of Particles 6 The Mole and Concentration 7 The Mole and Molar Volume 8 Empirical and Molecular Formulae 9 Chemical Equations 10 Calculations Based on Chemical Equations 11 Limiting Reagent 12 Gas Volume 13 Percent Yield Answers to Practice Problems Appendices **Supplemental Problems - Glencoe Chemistry Matter and Change** ,2013 Student's Solutions Manual for Introduction to Chemistry James Birk,Rich Bauer,Pamela Marks,2012-02-03 Each chapter of the Student Study Guide begins with a chapter review tied to the chapter goals in the text Next Sample problems are supplied and stepped out through the solution for each type of problem covered in the chapter A Self Test serves up fill in the blank exercises to assess learning with answers supplied at the end of the chapter Finally chapters end with the solutions for all of the in chapter problems as well as for the odd numbered end of chapter problems **3,000 Solved Problems In Chemistry** David Elliott Goldberg,1988 Master chemistry with Schaum s the high performance solved problem guide It will help you cut study time hone problem solving skills and achieve your personal best on exams Students love Schaum s Solved Problem Guides because they produce results Each year thousands of students improve their test scores and final grades with these indispensable guides Get the edge on your classmates Use Schaum s If you don t have a lot of time but want to excel in class use this book to Brush up before tests Study quickly and more effectively Learn the best strategies for solving tough problems in step by step detail Review what you ve learned in class by solving thousands of relevant problems that test your skill Compatible with any classroom text Schaum s Solved Problem Guides let you practice at your own pace and remind you of all the important problem solving techniques you need to remember fast And Schaum s are so complete they re perfect for preparing for graduate or professional exams Inside you will find 3000 solved problems with complete solutions the largest selection of solved problems yet published on this subject An index to help you quickly locate the types of problems you want to solve Problems like those you ll find on your exams Techniques for choosing the correct approach to problems Guidance toward the quickest most efficient solutions If you want

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