



COVALENT BONDING

Practice Problems

Do your answers to these very practice problems.

SECTION 8.1 MOLECULAR COMPOUNDS

1. Classify each of the following as ionic or a molecular compound.

- a. NaCl (ionic) b. CH_4 (molecular) c. HCl (molecular)
 d. H_2O (molecular) e. NaOH (ionic)

2. Which of the following is a molecular compound?

- a. NaCl b. H_2O c. HCl
 d. NaOH

3. Which of the following is a molecular compound? (Note: H_2O is a molecular compound.)

a. H_2O (molecular) b. NaCl (ionic) c. HCl (molecular) d. NaOH (ionic) e. H_2O (molecular)

4. Which of the following is a molecular compound? (Note: H_2O is a molecular compound.)

SECTION 8.2 THE NATURE OF COVALENT BONDING

1. Draw the Lewis structure for hydrogen fluoride, HF .



2. Draw the Lewis structure for methane, CH_4 .



3. Draw the Lewis structure for carbon dioxide, CO_2 .



4. Draw the Lewis structure for acetylene, C_2H_2 .

5. Draw the Lewis structure for hydrogen cyanide, HCN . (Note: $\text{H}-\text{C}\equiv\text{N}$ is the skeletal structure for hydrogen cyanide.)



Answer key:

SECTION 8.3 BONDING: THE ORBITAL

1. Which orbital has two lobes, one above and one below the nucleus? (a) s (b) p (c) d (d) f

2. Which orbital has two lobes, one to the left and one to the right of the nucleus? (a) s (b) p (c) d (d) f

Practice Problems Chapter 8 Covalent Bonding Worksheet

Robert R. Redfield



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