

Lab : Flame Test

Main Idea : Whoops!!! You're crazy teacher got her chemicals mixed up. It's your job to help her figure out which is which. This is what we know:

1. All of the chemicals are salts. A salt is any compound made of chloride (a halogen) and an element from groups 1 and 2 on the periodic table (Alkali & Alkaline Earth metals).
2. Some chemicals can be identified by the flame they produce. All of these compounds are **flammable**, meaning they burn. The different colors of the flames are the results of burning these different compounds.
3. The table to the right lists some chemicals and the colors they produce.

| Element | Color |
|-----------|--------------|
| Barium | Yellow-green |
| Calcium | Orange-red |
| Copper | Blue-green |
| Lithium | Crimson Red |
| Potassium | Pale Violet |
| Sodium | Yellow |
| Strontium | Red |

Materials : small beaker
paper towels
candle
4 wood splints
aluminum foil for candle wax
safety glasses for everyone
7 small vials of different salts

Safety First!

1. You must wear goggles at all times during the lab – these chemicals sizzle!
2. Do NOT touch the chemicals with your skin or clothing. If you do, wash immediately.
3. Keep hair tied back and sleeves rolled up.
4. Pay attention to what you are doing and to those around you.

Lab Procedures :

1. Read the following directions, #2 - 10, with your group before starting.
2. Fill your beaker with water and place your candle on the foil.
3. Raise your hand to begin. Your teacher will light the candle.
4. Choose one vial of "salt" from the collection. Carefully observe it with your **EYES ONLY**. Write down your observations of its physical properties on the chart.
5. **Predict the flame color** of this salt by recording your prediction on the chart.
6. Carefully have one person from your group dip the end (about 1/2 cm only) of the soaked wooden splint into the chemical so that you've collected a small amount of salt on its tip.
7. Carefully put the salt into the flame. For best results, aim the end of the stick to the side of the flame, rather than directly above it. Look very carefully, the flame will change color around the outside! It will not last long and you may only see it a little bit. **Observe and record the flame color.**
8. Dip the splint in the water. If it was not charred, you may be able to use again. Otherwise, use the other end of the splint for the next vial.
9. Repeat steps 4-8 for each vial of salt.

Clean Up : Throw away all used paper towels and splints. Make sure all vials are sealed tightly and placed back in the plastic bag. Place your candle and foil gently in your tub. Organize, clean counters, check bin back in to your teacher and return to your desk to work on the next few pages. Then, once settled, raise your hand to get a stamp →



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