Assessment.

National Contracts

## Section: Calculating Quantities in Reactions

In the space provided, write the letter of the term or phrase that best answers the question.

- The chemical equation P<sub>4</sub>O<sub>10</sub> + 6H<sub>2</sub>O → 4H<sub>3</sub>PO<sub>4</sub> can be correctly interpreted as
  - a. 1 molecule of P<sub>4</sub>O<sub>10</sub> + 6 molecules of H<sub>2</sub>O form 4 molecules of H-PO4.
  - b. 1 mol P<sub>4</sub>O<sub>10</sub> + 6 mol H<sub>2</sub>O reacts to form 4 mol H<sub>2</sub>PO<sub>4</sub>.
  - c. 1 g P<sub>4</sub>O<sub>10</sub> + 6 g H<sub>2</sub>O produces 4 g H<sub>3</sub>PO<sub>4</sub>.
  - d. Both (a) and (b)
- Stoichiometry is the branch of chemistry that deals with the between elements and compounds as reactants and products in a chemical reaction.
  - a. bonding
  - b. energy transfers
  - numerical relationship
  - physical interactions
- How many mole ratios can be correctly obtained from the following chemical equation?

$$Na_0S + Cd(NO_3)_2 \rightarrow 2NaNO_3 + CdS$$

- a. 1 b. 4
- c. 12
- d. 16
- 4. Molar mass can be used to relate
  - mass of reactant to moles of reactant.
  - b. moles of product to mass of product.
  - c. mass of product to moles of product.
  - d. All of the above
- In the chemical reaction described by the equation

$$4\text{Fe}(s) + 3\text{O}_2(g) \rightarrow 2\text{Fe}_2\text{O}_3(s)$$
,

the mole ratio of iron(III) oxide to iron is

- a. 1:1.
- b. 1-2.
- c. 2:3.
- d. 4:1.

# **Quiz Section Calculating Quantities In Reactions**

# **NARAYAN CHANGDER**

## **Quiz Section Calculating Quantities In Reactions:**

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