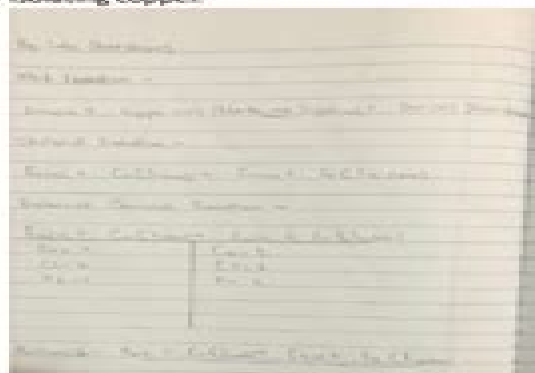


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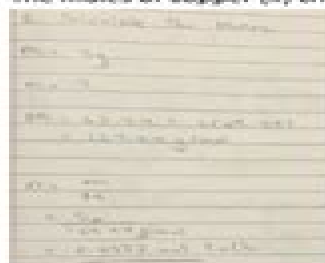
The percentage yield of a chemical reaction is determined by numerous factors: the nature of the reaction itself, the conditions under which the reaction was carried out and the nature of the reactants used. In this investigation, you will determine the limiting/excess reagents and the percentage yield of the reaction between iron and a solution of copper (II) chloride.

### Prelab Questions:

- Write the balanced chemical equation for the reaction between iron and copper (II) chloride. Remember to use your activity series to check the products.
  - The balanced equation is  $\text{Fe(s)} + \text{CuCl}_2\text{(aq)} \rightarrow \text{Cu(s)} + \text{FeCl}_2\text{(aq)}$ . The reaction is a single displacement isolating copper.



- Calculate the moles of copper (II) chloride in the solution if you use 5 g of copper (II) chloride.
  - The moles of copper (II) chloride in the solution, if I use 5g of copper (II) chloride, is 0.037 mol  $\text{CuCl}_2$ .



- Determine the limiting reagent if you use 1 g of iron.
  - The limiting reagent, if I use 1g of iron, is  $\text{Fe(s)}$ . It is 0.018mol because when comparing the numbers by the

## Limiting Reagents and % Yield



# Quick Lab Limiting Reagents Answers For 12 3

**Patrick Vollmar**



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