

Physical Science – 2nd Semester – Final Exam Study Guide – (2012–2013)

1. Physical Science Basics

- What tool(s) would you want to use to find the:
 - Mass of an object? triple beam balance Basic SI Unit? kg
 - Volume of an object? ruler or water displacement method Basic SI Unit? cm³ or ml
 - Length of an object? Ruler Basic SI Unit? meter
 - Density of an object? graduated cylinder Basic SI Unit? ml
- What is the formula for calculating density? mass/volume
 - What is the density of a cube that is 25 g and has a volume of 3 cm³? 8g/cm³
 - Compare the density of solid H₂O to liquid H₂O, and explain what observable effect this has. solid H₂O less dense than liquid H₂O

2. Chemistry Basics

- Element (def.) substance made up of one kind of atom
- Atom (def.) basic building block of matter
- Compound (def.) a molecule that contains atoms or molecules of two or more different elements
- Molecule (def.) a combination of two or more of the same or different atoms that are bonded together

3. Model of the Atomic Evolution of the Atomic Theory

Atoms are so small that many scientists throughout history have created models to describe them.

Why have atomic theories changed over time? Better technology came around.

Scientist	Year	Major Discovery/ Contribution	Name of atomic model
John Dalton	1808	Each element is made up of small, indivisible particles called atoms. Different elements have atoms of different masses.	Billiard Ball Model
J.J. Thomson	1897	Negatively charged electron	Plum Pudding Model
Niels Bohr	1913	Electrons move <u>fixed orbits</u> around the atom's nucleus in specific layers, or <u>electron shells</u> . Atoms absorb or give off energy when the electrons move from one shell (<u>energy level</u>) to another.	Bohr Model
Ernest Rutherford	1909	1909 /Based on idea that positive reflects positive/ atom mostly empty space; electrons orbit randomly around positively charged nucleus	Planetary Model

4. Periodic Table

- The current periodic table is arranged by increasing atomic number.
- Reading the Periodic Table
 - An element's properties can be predicted from its location in the periodic table.
 - The atomic number represents the number of protons.
 - The atomic mass represents the number of protons and neutrons.

Physical Science Final Exam Ii Study Guide

Annelies Wilder-Smith



Physical Science Final Exam Ii Study Guide:

Unveiling the Power of Verbal Beauty: An Psychological Sojourn through **Physical Science Final Exam Ii Study Guide**

In a global inundated with screens and the cacophony of immediate connection, the profound energy and psychological resonance of verbal artistry often diminish into obscurity, eclipsed by the regular barrage of noise and distractions. Yet, located within the lyrical pages of **Physical Science Final Exam Ii Study Guide**, a captivating function of literary elegance that impulses with organic feelings, lies an remarkable trip waiting to be embarked upon. Written by way of a virtuoso wordsmith, this interesting opus manuals visitors on a mental odyssey, lightly revealing the latent possible and profound influence embedded within the delicate web of language. Within the heart-wrenching expanse of the evocative examination, we will embark upon an introspective exploration of the book is key styles, dissect their fascinating publishing model, and immerse ourselves in the indelible impression it leaves upon the depths of readers souls.

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