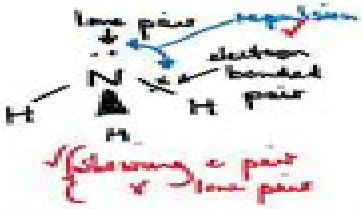
		gerangskik is om die hoeke tussen hulle te maksimeer. Strukturele elektronpare is bindingspare plus alleenpare. OF Valensdop-elektronpaarafstotingsmodel is 'n model om die vorms van molekules te voorspel waarin die elektronpaar vanaf die buitenste dop van 'n verwysingsatoom rondom hierdie atoom gerangskik is om die afstoting tussen hulle te minimaliseer. 3 marks/punte	
	1.2.4	The molecule is a tetrahedron Die molekule is 'n tetraeder 1 mark/punt	LO2
	1.2.5	 Name trigonal pyramid Noem trigonale piramidale 3 marks/punte	LO4
1.3	1.3.1	The electron pair is equally shared between the two Cl atoms and the resulting Cl_2 molecule has no degree of polarity. Die elektronpaar word ewe veel gedeel tussen die twee Cl-atome en die resulterende Cl2-molekule het geen mate van polariteit nie. 2 marks/punte	LO2
	1.3.2	The tendency of an atom to pull the bonded pair of electrons towards itself. Die neiging van 'n atoom om die gebinde elektronpare na homself te trek. 3 marks/punte	LO2
	1.3.3	Electronegativity of Na = 6 Electronegativity of Cl = 3 Difference is 3 With a difference of greater than 2.1 electron transfer will take place and the bond is ionic. (value of 2.1 does not need to be given for this mark) Elektronegatiwiteit van Na = 6 Elektronegatiwiteit van Cl = 3 Verskil is 3 Met 'n verskil van meer as 2,1 sal elektronoordrag plaasvind en die binding is ionies. (waarde van 2.1 hoef nie vir hierdie punt gegee te word nie) 4 marks/punte	LO2

Physical Science 2014 Control Test

VM Jensen

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