

## Quiz 1: Equilibria, Solids & Gases

### Multiple Choice

Answer the lower of the choice that best completes the statement or answers the question.

#### CONTENT REVIEW

- \_\_\_\_\_ 1. Chemical equilibrium exists when
- the processes cease.
  - the driving forward processes cease.
  - the opposing processes occur at different rates.
  - the opposing processes occur at the same rate.
- \_\_\_\_\_ 2. Any chemical reaction in which the products are represented by brackets is called a(n)
- double-replacement reaction.
  - single-replacement reaction.
  - reversible reaction.
  - precipitation reaction.
- \_\_\_\_\_ 3. What symbol is used to indicate a reversible reaction?
- $\rightleftharpoons$
  - $\rightleftharpoons$
  - $\rightleftharpoons$
  - $\rightleftharpoons$
- \_\_\_\_\_ 4. What is the general relationship between the concentrations and the rate of a forward reaction?
- Rate is greater when product concentrations is greater.
  - Rate is greater when reactant concentrations is greater.
  - Rate is greater when catalyst concentrations is greater.
  - No relationship exists between concentrations and rate.
- \_\_\_\_\_ 5. What symbol is used to denote concentration?
- $[ ]$
  - $( )$
  - $\{ \}$
  - $\%$
- \_\_\_\_\_ 6. At chemical equilibrium, which of the following is always true of the concentrations of reactants and products?
- They are constant.
  - They are equal.
  - Reactant concentrations is zero.
  - Product concentrations is zero.
- \_\_\_\_\_ 7. What is the equilibrium constant for the general equation  $aA + bB \rightleftharpoons cC + dD$ ?
- $\frac{[A]^a[B]^b}{[C]^c[D]^d}$
  - $\frac{[C]^c[D]^d}{[A]^a[B]^b}$
  - $\frac{[C]^c[D]^d}{[A]^a[B]^b}$
  - $\frac{[C]^c[D]^d}{[A]^a[B]^b}$
- \_\_\_\_\_ 8. The symbol for the equilibrium constant is
- $K'$
  - $K_{eq}$
  - $K_{eq}$
  - $K'$
- \_\_\_\_\_ 9. The principle that relates changes in pressure on equilibrium systems to equilibrium position is
- Boyle's law.
  - Le Chatelier's principle.
  - the law of chemical equilibrium.
  - Avogadro's principle.
- \_\_\_\_\_ 10. If a change in conditions is imposed on a system at equilibrium, the equilibrium position will always

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