

## EQUILIBRIUM CONSTANT - WORKSHEET 2\_ANSWERS

1. Write the equilibrium constant expression for each of the following reactions:



$$K = \frac{[\text{NO}_2]^2}{[\text{N}_2][\text{O}_2]}$$



$$K = \frac{[\text{C}_2\text{H}_6]}{[\text{H}_2]^3}$$



$$K = \frac{1}{[\text{I}_2(\text{aq})]}$$

2. For the following reaction:



At equilibrium,  $[\text{Fe}^{2+}] = 0.50\text{ M}$ ,  $[\text{Ag}^+] = 1.0\text{ M}$ ,  $[\text{Fe}^{3+}] = 1.50\text{ M}$

- a) Calculate the equilibrium constant for the reaction.

$$K = 3.0$$

- b) If at equilibrium,  $[\text{Fe}^{2+}] = 0.20\text{ M}$  and  $[\text{Ag}^+] = 0.30\text{ M}$ , what must be the concentration of  $\text{Fe}^{3+}$ ?

$$[\text{Fe}^{3+}] = 0.18\text{ mol/L}$$

3.  $2\text{NO}(\text{g}) + \text{O}_2(\text{g}) \rightleftharpoons 2\text{NO}_2(\text{g}) \quad K = 5.0 \times 10^{-6}$

Calculate the equilibrium concentration of oxygen gas if the NO and  $\text{NO}_2$  concentrations are equal at equilibrium.

$$[\text{O}_2] = 2.0 \times 10^{-6}\text{ mol/L}$$

4. a)  $\text{CH}_3\text{COOH}(\text{aq}) \rightleftharpoons \text{H}^+(\text{aq}) + \text{CH}_3\text{COO}^-(\text{aq}) \quad K = 1.8 \times 10^{-5}$



Which of the above reactions would have the greatest amount of products present at equilibrium compared to reactants?

Reaction b. Since K is larger the amount of products must be larger, because K is products/reactants.

5. For the reaction:



at an equilibrium temperature of  $1000^\circ\text{C}$ , a 1.00 L flask contains 0.120 moles of ammonia, 1.03 moles of nitrogen and 1.62 moles of hydrogen. Calculate the equilibrium constant at this temperature.

$$K = 3.29 \times 10^{-2}$$

# Questions On Equilibrium Constants And Answers

**Christian Drosten**



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## **Table of Contents Questions On Equilibrium Constants And Answers**

1. Understanding the eBook Questions On Equilibrium Constants And Answers
  - The Rise of Digital Reading Questions On Equilibrium Constants And Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Questions On Equilibrium Constants And Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Questions On Equilibrium Constants And Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Questions On Equilibrium Constants And Answers
  - Personalized Recommendations
  - Questions On Equilibrium Constants And Answers User Reviews and Ratings
  - Questions On Equilibrium Constants And Answers and Bestseller Lists
5. Accessing Questions On Equilibrium Constants And Answers Free and Paid eBooks

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  - Questions On Equilibrium Constants And Answers Budget-Friendly Options
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    - ePub, PDF, MOBI, and More
    - Questions On Equilibrium Constants And Answers Compatibility with Devices
    - Questions On Equilibrium Constants And Answers Enhanced eBook Features
  7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Questions On Equilibrium Constants And Answers
    - Highlighting and Note-Taking Questions On Equilibrium Constants And Answers
    - Interactive Elements Questions On Equilibrium Constants And Answers
  8. Staying Engaged with Questions On Equilibrium Constants And Answers
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Questions On Equilibrium Constants And Answers
  9. Balancing eBooks and Physical Books Questions On Equilibrium Constants And Answers
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Questions On Equilibrium Constants And Answers
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Questions On Equilibrium Constants And Answers
    - Setting Reading Goals Questions On Equilibrium Constants And Answers
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Questions On Equilibrium Constants And Answers
    - Fact-Checking eBook Content of Questions On Equilibrium Constants And Answers
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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