#### Honors Chemistry Quiz Chapter 6: Thermochemistry

Name: \_\_\_\_\_ Period: \_\_\_\_ Date: \_\_\_\_ This quiz is worth 40 points; each correct response is 2 points. Only those quizzes completed in black ink will be graded. Good luck!

 Energy that is associated with the position or composition of an object is called

A) kinetic energy

B) thermal energy

C) potential energy

D) chemical energy

E) None of the above.

2) Which of the following signs on q and w represent a system that is doing work on the surroundings, as well as losing heat to the surroundings?

A) q = -, w = -

B) q = +, w = + C) q = -, w = +

D) q = +, w = -

E) None of these represent the system referenced above.

Calculate the change internal energy (ΔE) for a system that is giving off
45.0 kJ of heat and is performing 855 J of work on the surroundings.

A) 44.1 kJ B) -44.1 kJ

C) -45.9 kJ

D) 9.00 x 102 kJ

E) -9.00 x 102 kJ

substance by 1°C

4) Define specific heat capacity.

A) the quantity of heat required to raise the temperature of 1 mole of a substance by 1°C

B) the quantity of heat required to change a system's temperature by 1°C C) the quantity of heat required to raise the temperature of 1 gram of a

D) the quantity of heat required to raise the temperature of 1 gram of a substance by 1°F

E) the quantity of heat required to raise the temperature of 1 liter of a substance by 1°C

# **Questions For Honors On Thermochemistry**

**The Princeton Review** 

#### **Questions For Honors On Thermochemistry:**

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