

# Periodic Trends Worksheet

Directions: Use the trends in electronegativity to answer the questions.

1. Write the following elements by increasing atomic radius: sodium, sulfur, chlorine, oxygen, potassium.

**K, Na, S, Cl, O**

2. Write the following elements by increasing electronegativity: sodium, oxygen, sulfur, chlorine, potassium.

**O, Cl, S, Na, K**

3. Which atom has a higher electronegativity: sodium or sulfur?

**Sulfur (sulfur is to the right of sodium and closer to fluorine)**

4. Which two elements in the same family generally have similar properties?

**There have been two cases here where a strong periodic trend is present.**

5. Evaluate whether the following properties increase or decrease from left to right across the periodic table:

- a. atomic radius (decreasing to the right) **decreases**
- b. first ionization energy **increases**
- c. electronegativity **increases**

6. Which period in the periodic table shows atoms that are (a) periodic table? (b) the elements, this trend?

**Periods 2 and 3 show strong trends in electronegativity, atomic radius, and ionization energy.**

7. Which trend in ionization energy across atoms is predicted and the periodic table? Which trend is unexpected?

**Expected: increasing electronegativity from left to right. Unexpected: decreasing ionization energy.**

8. Write the atoms in each group that have the highest and lowest values.

|    |    |
|----|----|
| Li | Be |
| Na | Mg |
| K  | Ca |
| Rb | Sr |
| Cs | Ba |
| Fr | Ra |

|    |    |    |    |   |    |    |
|----|----|----|----|---|----|----|
| Li | Be | B  | C  | N | O  | F  |
| Na | Mg | Al | Si | P | S  | Cl |
| K  | Ca | Sc | Ti | V | Cr | Mn |

9. Circle the atoms in each group that have the highest electronegativity.

# Physical Science If8767 Pg 63 Key

**BM King**



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