

QUESTION 10

The three test tubes, X, Y and Z below, contain colourless solutions. The solutions in each test tube can either be potassium iodide or sodium carbonate or magnesium sulphate. Tests are carried out to determine which test tube contains which salts.



The following tests were carried out and the observations were made:

- Barium chloride solution is added to solution X; the solution remains colourless.
- Barium chloride solution is added to Y and Z; a white precipitate is formed in both cases.
- Dilute nitric acid is added to the precipitate in solution Y; the precipitate dissolves.

(2)

10.1 Define the term *dissociation*.

10.2 Give a reason why barium chloride solution can be used as electrolyte.

(2)

10.3 Write down the name of the precipitate in test tube Z.

(1)

Use the above information to identify the solutions in each of the test tubes X, Y and Z.

10.4 Write down the LETTER that represents the test tube and next to it, the NAME of the solution and the reason for the choice.

(6)

10.5 What type of reaction takes place between the precipitate in Y and the nitric acid?

(1)

[12]

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