

Name: _____

CLASSIFYING CHEMICAL REACTIONS

Classify the following chemical reactions as one of the following: synthesis, decomposition, single-replacement, double-replacement, or combustion.

$2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$	$\text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$
$2\text{K} + 2\text{H}_2\text{O} \rightarrow \text{H}_2 + 2\text{KOH}$	$\text{CaCO}_3 + 2\text{HCl} \rightarrow \text{CaCl}_2 + \text{H}_2\text{CO}_3$
$2\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$	$\text{Ca} + 2\text{HCl} \rightarrow \text{CaCl}_2 + \text{H}_2$
$2\text{SO}_2 + \text{O}_2 \rightarrow 2\text{SO}_3$	$\text{Pb(NO}_3)_2 + 2\text{HCl} \rightarrow \text{PbCl}_2 + 2\text{HNO}_3$
$\text{Cu} + 2\text{AgNO}_3 \rightarrow 2\text{Ag} + \text{Cu(NO}_3)_2$	$2\text{Na} + \text{Cl}_2 \rightarrow 2\text{NaCl}$
$2\text{H}_2\text{O} \rightarrow 2\text{H}_2 + \text{O}_2$	$\text{CH}_4 + 2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$
$\text{Pb(NO}_3)_2 + 2\text{K} \rightarrow \text{Pb} + 2\text{KNO}_3$	$\text{H}_2\text{CO}_3 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
$2\text{Fe} + \text{O}_2 \rightarrow 2\text{FeO}$	$\text{Zn} + 2\text{HCl} \rightarrow \text{H}_2 + \text{ZnCl}_2$
$\text{HCl} + \text{NaOH} \rightarrow \text{NaCl} + \text{H}_2\text{O}$	$2\text{Na} + 2\text{H}_2\text{O} \rightarrow 2\text{NaOH} + \text{H}_2$

Physical Science If8767 Classifying Chemical Reactions

Michal Rosen-Zvi



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web the set of polynomials is closed under addition and subtraction so the sum or difference of any two polynomials is also a polynomial to add polynomials add like terms you can use a vertical or a horizontal format adding polynomials find the sum a $2x^3 - 5x^2 - x - 2x^2 - x^3 - 1$ b $3x^2 - x - 6x^2 - 4x - 10$ solution a

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n3 1 3n3 2 6a 3a2 2a2 3a 3 x2 x 8x 2x2 4 2a2 4a3 3a3 8 5 5x2 4 5 5x3 6 8n2 2n3 6n3 8n2

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classify by number of terms 7x 3 8x 2 4x 9 monomial

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or subtracting their numerical coefficients however the exponents of like terms should not be added or subtracted as they
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