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web 11 use avogadro's law to determine the number of moles of butane gas assume that butane is an ideal gas and that one mole has a volume of 22.4 l at STP $V_1/n_1 = V_2/n_2$ $0.0896 \text{ l} / n = 22.4 \text{ l} / 1 \text{ mole}$ $n = 0.0040 \text{ moles}$ 12 experimental molar mass g mole of butane 0.23 g / 0.0040 moles = 58 g mole 13 the molecular formula of butane is C_4H_{10}

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