

PSYCHOLOGY 2800W

SPRING 2013

TENTATIVE SYLLABUS; CAN BE CHANGED AT ANY TIME!!

Psychology 2800W: Social Psychology Spring, 2013

Instructor: Cassin Ganevy, M.A.

Office Hours: W 9:00-12:00, 223 Science I

Dates: 01/14/2013 to 03/08/2013 (first half)

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Phone: 515-254-6507 (office)

Instructor: Yuenan Yang, M.S.

Office Hours: F 2-5 PM, 223 Science I

Dates: 03/11/2013 to 05/03/2013 (second half)

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COURSE OVERVIEW

ISU Catalogue description: "Individual human behavior in social contexts. Emphasis on social judgments and decisions, attitudes, perceptions of others, social influence, aggression, stereotypes, and helping."

Course goals: By the end of this course, you should:

- Have an increased knowledge base regarding social psychology, as well as an enhanced ability to examine social behavior with a critical eye
- Be able to explain how psychologists do what they do, and why they do it that way
- Be able to recognize and apply various social psychological principles to your everyday life

Required textbook: *Social Psychology* (2007, 7th edition) by Kassir, Fein, & Markus.

Course format: This course will be delivered in an online format.

Course materials: Course notes, audio lectures, and other materials will be available on Blackboard Learn.

STUDENT EVALUATION

Journal discussion board postings: Because this is an online-only course, there are no chances for face-to-face contact, which eliminates one of the most important aspects of a course – the opportunity to ask questions and to participate in discussion. To foster a "classroom" environment, each student will be required to contribute to the class via discussion boards. Students in previous classes have really enjoyed this assignment.

Students will be required to post a journal on the discussion boards of Blackboard Learn. Each journal post will count for 8 points. Please see the "Social Psychology Journal" file on Blackboard Learn for a more detailed description of what is expected, as well as several example entries. We encourage you to check the discussion boards often.

Journal responses: Students are required to make 3 responses/comments on journals posted by classmates. Each comment you make will count for 1 point. Comments must be thoughtful – something like, "Oh yeah, dude, that's cool!" will not count. You could post something about a similar or different experience, or how you agree with the original poster and why, or questions you might have for the original poster. Comments are due before the next journal is due.

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web these ions are called spectator ions because they do not participate in the actual reaction canceling the spectator ions gives the net ionic equation which shows only those species that participate in the chemical reaction $2\text{Ag}^+ \text{aq} + \text{Cr}_2\text{O}_7^{2-} \text{aq} \rightarrow 2\text{Cr}_2\text{O}_7^{2-} \text{aq} + 2\text{Ag}^+ \text{aq}$

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web chemical equation in which all dissolved ionic reactants and products including spectator ions are explicitly represented by formulas for their dissociated ions molecular equation chemical equation in which all reactants and products are

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web jan 29 2020 the ionic form of this reaction is $2\text{Na}^+ \text{aq} + 2\text{Cl}^- \text{aq} + \text{Cu}^{2+} \text{aq} + \text{SO}_4^{2-} \text{aq} \rightarrow 2\text{Na}^+ \text{aq} + \text{SO}_4^{2-} \text{aq} + \text{CuCl}_2 \text{s}$ the sodium ions and sulfate ion are the spectator ions in this reaction they appear unchanged in both the product and reactant side of the equation

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web mar 11 2014 solution ionic equation $\text{H}^+ \text{aq} + \text{Cl}^- \text{aq} + \text{Na}^+ \text{aq} + \text{HCO}_3^- \text{aq} \rightarrow \text{Na}^+ \text{aq} + \text{Cl}^- \text{aq} + \text{H}_2\text{O} \text{l} + \text{CO}_2 \text{g}$ the spectator ions are Cl^- and Na^+ net ionic equation $\text{H}^+ \text{aq} + \text{HCO}_3^- \text{aq} \rightarrow \text{H}_2\text{O} \text{l} + \text{CO}_2 \text{g}$ here is a video on spectator ions [chapter 8 04 ionic equations chemistry libretexts](#) - Mar 06 2023

web the six $\text{NO}_3^- \text{aq}$ ions and the six $\text{Na}^+ \text{aq}$ ions that appear on both sides of the equation are spectator ions that can be canceled to give the net ionic equation $3\text{Ba}^{2+} \text{aq} + 2\text{PO}_4^{3-} \text{aq} \rightarrow \text{Ba}_3(\text{PO}_4)_2 \text{s}$

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web these spectator ions ions whose presence is required to maintain charge neutrality are neither chemically nor physically changed by the process and so they may be eliminated from the equation to yield a more succinct representation called a

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web practice problems on net ionic equations page 2 of 3 answer key to practice problems on net ionic equations 1 molecular $\text{AgNO}_3 \text{aq} + \text{KCl} \text{aq} \rightarrow \text{AgCl} \text{s} + \text{KNO}_3 \text{aq}$ total ionic $\text{Ag}^+ \text{aq} + \text{NO}_3^- \text{aq} + \text{K}^+ \text{aq} + \text{Cl}^- \text{aq} \rightarrow \text{AgCl} \text{s} + \text{K}^+ \text{aq} + \text{NO}_3^- \text{aq}$ net ionic $\text{Ag}^+ \text{aq} + \text{Cl}^- \text{aq} \rightarrow \text{AgCl} \text{s}$

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web video answer you know spectator ions are the ions present on the both sides of the equation indicate cancels out in the total ionic equations so you know in a case chemical reactions are described with chemical equations which symbolically present the starting materials end products

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web net ionic equations an equation for a reaction in solution showing only those particles that are directly involved in the chemical change spectator ions an ion that is not directly involved in a chemical reaction an ion that does not change

oxidation number or composition during a reaction products

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web the equation below is the final net ionic equation after elimination of the spectator ions which yields water h aq oh aq h₂ o l the spectator ions in koh aq and hno₃ aq are k aq and no₃ aq how to identify and list the

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web the ag aq and cl aq ions become agcl s but the na aq ions and the no₃ aq ions stay as na aq ions and no₃ aq ions these two ions are examples of spectator ions ions that do nothing in the overall course of a chemical reaction

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web aug 8 2022 a spectator ion is an ion that does not take part in the chemical reaction and is found in solution both before and after the reaction the net ionic equation is the chemical equation that shows only those elements compounds and ions that are directly involved in the chemical reaction

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web jun 2 2023 the spectator ions are cut out from a reaction hence the net ionic equation for the above reaction would be $2\text{Cl}^- + \text{Cu}^{2+} + 2\text{Cl}^- + \text{Cu}^{2+} \rightarrow 2\text{Cl}^- + \text{Cu}^{2+} + 2\text{Cl}^- + \text{Cu}^{2+}$ these spectator ions are not actively involved in the reaction the existence of spectator ions affects the debye length

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