Chapter 13 Chemical Kinetics

Student: __

- 1. The units of "reaction rate" are
 - A. L mol⁻¹ s⁻¹.
 - B. L² mol⁻² s⁻¹.
 - C. 8
 - E mol L st.
- For the reaction BrO₃" + 5Br" + 6H" → 3Br₂ + 3H₂O at a particular time, _ΔΔ[BrO₃"]/Δt = 1.5 × 10⁻² M/s. What is -Δ[Br"]/Δt at the same instant?
 - A. 13 M/s
 - B. $7.5 \times 10^{-3} \text{ M/s}$
 - C 1.5 × 10°2 M/s
 - $3.0 \times 10^{-3} \text{ M/s}$
 - E. 330 M/s
- For the following reaction, ΔP(C₆H₁₄)/ΔI was found to be -6.2 × 10⁻³ atm/s. C₆H₁₄(g) → C₆H₆(g) + 4H₂(g)

Determine $\Delta P(H_2)/\Delta t$ for this reaction of the same time.

- A. 6.2×10^{13} atm/s
- B. 1.6 × 10⁻³ atm/s
- C. $2.5 \times 10^{-2} \text{ atm/s}$
- D. -1.6 × 10⁻³ atm/s
- E. -2.5 × 10⁻² ktm/s
- For the reaction C₆H₁₄(g) → C₆H₆(g) + 4H₂(g), ΔP(H₂)/Δt was found to be 2.5 × 10⁻² atm/s, where ΔP(H₂) is the change in pressure of hydrogen. Determine ΔP(C₆H₁₄)/Δt for this reaction at the same time.
 - A. 2.5 × 10⁻² atm/s
 - B. -6.2×10^{-3} atm/s
 - C. -2.5×10^{-2} atm/s
 - D. 0.10 atm/s
 - E. $6.2 \times 10^{-3} \text{ atm/s}$

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