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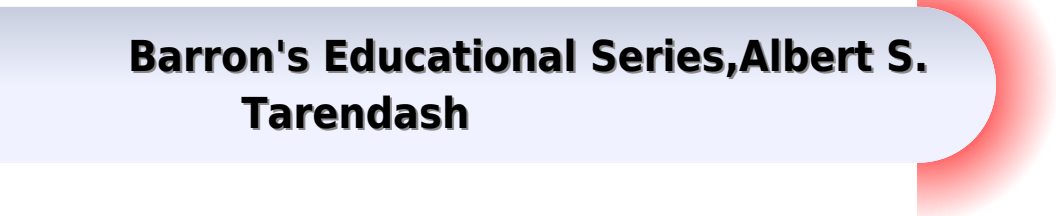
Chemistry Practice: Writing Chemical Formulas

Write a chemical formula for each substance.

- | | | | |
|--|--------------------------------|--|---------------------------|
| 1. <u>NaCl</u> | sodium chloride | 39. <u>Ca(NO₃)₂</u> | calcium nitrate |
| 2. <u>N₂O₅</u> | dinitrogen pentoxide | 40. <u>SO₃</u> | sulfur trioxide |
| 3. <u>H₂S</u> | hydrosulfuric acid | 41. <u>KCN</u> | potassium cyanide |
| 4. <u>K₂SO₄</u> | potassium sulfate | 42. <u>Pb(NO₃)₂</u> | lead(II) nitrate |
| 5. <u>H₂C₂O₄</u> | oxalic acid | 43. <u>H₂S</u> | hydrogen sulfide |
| 6. <u>Ag₂C₂H₃O₂</u> | silver acetate | 44. <u>CoCl₃</u> | cobalt(III) chloride |
| 7. <u>Cr(NO₃)₃</u> | chromium(III) nitrate | 45. <u>SF₆</u> | sulfur hexafluoride |
| 8. <u>HC₂H₃O₂</u> | acetic acid | 46. <u>Ca₃N₂</u> | calcium nitride |
| 9. <u>(NH₄)₂CO₃</u> | ammonium carbonate | 47. <u>CuI</u> | copper(I) iodide |
| 10. <u>Ca(OH)₂</u> | calcium hydroxide | 48. <u>SiO₂</u> | silicon dioxide |
| 11. <u>H₂C₄H₄O₆</u> | tartaric acid | 49. <u>Sn(C₂H₃O₂)₄</u> | tin(IV) acetate |
| 12. <u>Hg(NO₃)₂</u> | mercury(II) nitrate | 50. <u>CCl₄</u> | carbon tetrachloride |
| 13. <u>N₂O</u> | dinitrogen monoxide | 51. <u>CuS</u> | copper(II) sulfide |
| 14. <u>Fe₂O₃</u> | iron(III) oxide | 52. <u>Pb₃(PO₄)₂</u> | lead(II) phosphate |
| 15. <u>Pb(ClO₃)₂</u> | lead(II) chlorate | 53. <u>XeCl₄</u> | xenon tetrachloride |
| 16. <u>(NH₄)₃PO₄</u> | ammonium phosphate | 54. <u>Rb₂O</u> | rubidium oxide |
| 17. <u>ZnCl₂</u> | zinc chloride | 55. <u>MgSe</u> | magnesium selenide |
| 18. <u>Ca₃(PO₄)₂</u> | calcium phosphate | 56. <u>NH₄Cl</u> | ammonium chloride |
| 19. <u>OF₂</u> | oxygen difluoride | 57. <u>Fe(C₂H₃O₂)₃</u> | iron(III) acetate |
| 20. <u>NaN₃</u> | sodium azide | 58. <u>K₂Cr₂O₇</u> | potassium dichromate |
| 21. <u>Fe₂(SO₄)₃</u> | iron(III) sulfate | 59. <u>PBr₃</u> | phosphorous tribromide |
| 22. <u>H₃AsO₃</u> | arsenous acid | 60. <u>Na₃PO₃</u> | sodium phosphite |
| 23. <u>Cr₂O₃</u> | chromium(III) oxide | 61. <u>Na₃PO₄</u> | sodium phosphate |
| 24. <u>N₂O₄</u> | dinitrogen tetroxide | 62. <u>Hg(NO₃)₂</u> | mercury(II) nitrate |
| 25. <u>NH₄NO₃</u> | ammonium nitrate | 63. <u>LiHCO₃</u> | lithium bicarbonate |
| 26. <u>AsBr₃</u> | gold(III) bromide | 64. <u>CrF₃</u> | chromium(III) fluoride |
| 27. <u>CO</u> | carbon monoxide | 65. <u>PbI₂</u> | lead(II) iodide |
| 28. <u>K₂CO₃</u> | potassium carbonate | 66. <u>H₂SO₃</u> | sulfurous acid |
| 29. <u>HIO₃</u> | iodic acid | 67. <u>SnF₂</u> | tin(II) fluoride |
| 30. <u>CsCl</u> | cesium chloride | 68. <u>HgCrO₄</u> | mercury(II) chromate |
| 31. <u>Ni(MnO₄)₂</u> | nickel(II) permanganate | 69. <u>KNO₃</u> | potassium nitrate |
| 32. <u>Al₂(SO₄)₃</u> | aluminum sulfate | 70. <u>SrCl₂</u> | strontium chloride |
| 33. <u>Al₂(SO₃)₃</u> | aluminum sulfite | 71. <u>P₄O₁₀</u> | tetraphosphorous decoxide |
| 34. <u>Ba(C₂H₃O₂)₂</u> | barium acetate | 72. <u>KNO₃</u> | potassium nitrate |
| 35. <u>Mn(OH)₃</u> | manganese(III) hydroxide | 73. <u>KNO₂</u> | potassium nitrite |
| 36. <u>KH₂PO₄</u> | potassium dihydrogen phosphate | 74. <u>K₃N</u> | potassium nitride |
| 37. <u>HF</u> | hydrofluoric acid | 75. <u>CaO</u> | calcium oxide |
| 38. <u>Al(BrO₃)₃</u> | aluminum bromate | 76. <u>Fe(IO₄)₂</u> | iron(II) periodate |

Nomenclature And Formula Lab Answer Key

**Barron's Educational Series, Albert S.
Tarendash**



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from this to determine its hysteresis loss 17 To determine the ballistic constant of a moving coil ballistic galvanometer 18 To determine the coefficient of viscosity of water by Poiseuille's method 19 To determine the coefficient of viscosity of a liquid by rotating viscometer 20 To measure fiber attenuation and numerical aperture of fiber 21 To determine high resistance by leakage method 22 To determine magnetic susceptibility of a paramagnetic solution by Quincke's method **Instructors**

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