

## Quantum Number Problems

Use the rules for quantum numbers and notes to answer the questions below:

Rules: (1)  $n = 1, 2, \dots, \infty$

(2)  $l = 0, 1, 2, 3, \dots, n-1$  or s ( $l = 0$ ), p ( $l = 1$ ), d ( $l = 2$ ), f ( $l = 3$ )

(3)  $m_l = -l, -(l-1), \dots, 0, \dots, +(l-1), +l$  (e.g., when  $l = 1$ ,  $m_l = -1, 0, +1$ )

(4) The number of nodal planes ( $N_p$ ) is equal to the angular momentum quantum number. This means  $l = N_p$ .

(5) The total number of nodes is  $N_T = N_p + N_r$ , where  $N_r$  is the number of spherical nodes (radial nodes).

$N_T = n - 1$  where  $n$  is the principal quantum number. So  $n = N_p + N_r + 1$

From these orbitals:    2s    3d    4p    5f    6d    7p

Fill in the blanks with the correct answer(s):

*(answers may be used more than once or not at all; none is a possible answer)*

- |                                            |       |                                           |       |
|--------------------------------------------|-------|-------------------------------------------|-------|
| 1. The orbital(s) with $n = 6$             | _____ | 4. The orbital(s) with 0 planar nodes     | _____ |
| 2. The orbital(s) with $l = 1$             | _____ | 5. The orbital(s) with 3 spherical nodes  | _____ |
| 3. The orbital(s) that can have $m_l = -2$ | _____ | 6. The orbital(s) with a total of 5 nodes | _____ |

# Problems On Quantum Numbers

**Thomas Griffiths**



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