

GAS LAWS OVERVIEW

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Background: The gas laws we will be discussing in this handout were created over four centuries ago, and have been helpful for scientists to find pressures, amounts, volumes and temperatures of gases under different conditions, and the relationship there exists between these variables.

The fundamental gas laws are the following: **Boyle's Law**, **Charles' Law**, and **Avogadro's Law**. We will also discuss the **Gay-Lussac law**. When we combine these Laws, we get the **Combined Gas Law** and the **Ideal Gas Law**.

GAS LAWS:

- **Boyle's Law:** In 1662, Robert Boyle discovered the relationship between **pressure** and **volume**, assuming constant temperature and amount of gas. If the temperature remains constant, as the volume of a container increases, the pressure the gas will exert will decrease. Therefore:

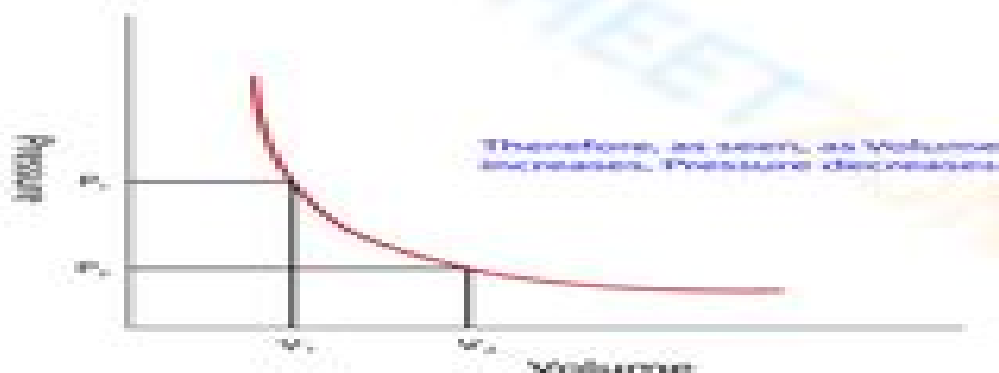
RELATIONSHIP:

Pressure is inversely proportional to volume

EQUATION:

$$P_1 \times V_1 = P_2 \times V_2$$

GRAPHIC REPRESENTATION:



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