

Electron Configuration Practice - Homework - **KEY**

In the space below, write the expanded electron configurations of the following elements:

- | | | |
|----|-----------|--|
| 1) | sodium | $1s^2 2s^2 2p^6 3s^1$ |
| 2) | potassium | $1s^2 2s^2 2p^6 3s^2 3p^6 4s^1$ |
| 3) | chlorine | $1s^2 2s^2 2p^6 3s^2 3p^5$ |
| 4) | bromine | $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^5$ |
| 5) | oxygen | $1s^2 2s^2 2p^4$ |

In the space below, write the abbreviated electron configurations of the following elements:

- | | | |
|-----|-----------|-------------------------|
| 6) | manganese | $[\text{Ne}] 3s^2$ |
| 7) | silver | $[\text{Kr}] 5s^2 4d^9$ |
| 8) | nitrogen | $[\text{He}] 2s^2 2p^3$ |
| 9) | sulfur | $[\text{Ne}] 3s^2 3p^4$ |
| 10) | argon | $[\text{Ne}] 3s^2 3p^6$ |

In the space below, write the orbital notation of the following elements:

- | | | | | | | | | | | | | | | | | | |
|---------------|-------------------|----------------------------|----------------------------|----------------------------|----------------------------|----------------------------|----------------------------|----------------------------|----------------------------|----------------------------|----------------------------|----------------------------|----------------------------|----------------------------|----------------------------|----------------------------|------------------|
| 11) manganese | 12 e ⁻ | $\uparrow\downarrow$
1s | $\uparrow\downarrow$
2s | $\uparrow\downarrow$
2p | $\uparrow\downarrow$
2p | $\uparrow\downarrow$
2p | $\uparrow\downarrow$
3s | | | | | | | | | | |
| 12) gallium | 31 e ⁻ | $\uparrow\downarrow$
1s | $\uparrow\downarrow$
2s | $\uparrow\downarrow$
2p | $\uparrow\downarrow$
2p | $\uparrow\downarrow$
2p | $\uparrow\downarrow$
3s | $\uparrow\downarrow$
3p | $\uparrow\downarrow$
3p | $\uparrow\downarrow$
3p | $\uparrow\downarrow$
4s | $\uparrow\downarrow$
3d | $\uparrow\downarrow$
3d | $\uparrow\downarrow$
3d | $\uparrow\downarrow$
3d | $\uparrow\downarrow$
3d | $\uparrow$
4p |
| 13) nitrogen | 7 e ⁻ | $\uparrow\downarrow$
1s | $\uparrow\downarrow$
2s | $\uparrow$
2p | $\uparrow$
2p | $\uparrow$
2p | | | | | | | | | | | |
| 14) sulfur | 16 e ⁻ | $\uparrow\downarrow$
1s | $\uparrow\downarrow$
2s | $\uparrow\downarrow$
2p | $\uparrow\downarrow$
2p | $\uparrow\downarrow$
2p | $\uparrow\downarrow$
3s | $\uparrow\downarrow$
3p | $\uparrow$
3p | $\uparrow$
3p | | | | | | | |
| 15) argon | 18 e ⁻ | $\uparrow\downarrow$
1s | $\uparrow\downarrow$
2s | $\uparrow\downarrow$
2p | $\uparrow\downarrow$
2p | $\uparrow\downarrow$
2p | $\uparrow\downarrow$
3s | $\uparrow\downarrow$
3p | $\uparrow\downarrow$
3p | $\uparrow\downarrow$
3p | | | | | | | |

Determine what elements are denoted by the following electron configurations:

- | | | |
|-----|---|-------------|
| 16) | $1s^2 2s^2 2p^6 3s^2 3p^4$ | sulfur |
| 17) | $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^1$ | rubidium |
| 18) | $[\text{Kr}] 5s^2 4d^{10} 5p^3$ | antimony |
| 19) | $[\text{Xe}] 6s^2 4f^{14} 5d^6$ | osmium |
| 20) | $[\text{Rn}] 7s^2 5f^{11}$ | einsteinium |

Determine which of the following electron configurations are not valid:

- | | | |
|-----|--|--|
| 21) | $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 4d^{10} 4p^5$ | not valid (take a look at "4d") |
| 22) | $1s^2 2s^2 2p^6 3s^3 3d^5$ | not valid (3p comes after 3s) |
| 23) | $[\text{Ra}] 7s^2 5f^6$ | not valid (radium isn't a noble gas) |
| 24) | $[\text{Kr}] 5s^2 4d^{10} 5p^5$ | valid |
| 25) | $[\text{Xe}]$ | not valid (an element can't be its own electron configuration) |

Science Electron Configuration Packet Answer Key

Ralph Thomas Myers



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A Self-study Guide to the Principles of Organic Chemistry Jiben Roy, 2013 A Self Study Guide to the Principles of Organic Chemistry Key Concepts Reaction Mechanisms and Practice Questions for the Beginner will help students new to organic chemistry grasp the key concepts of the subject quickly and easily as well as build a strong foundation for future study Starting with the definition of atom the author explains molecules electronic configuration bonding hydrocarbons polar reaction mechanisms stereochemistry reaction varieties organic spectroscopy aromaticity and aromatic reactions biomolecules organic polymers and a synthetic approach to organic compounds The over one hundred diagrams and charts contained in this volume will help students visualize the structures and bonds as they read the text and make the logic of organic chemistry clear and easily understood Each chapter ends with a list of frequently asked questions and answers followed by additional practice problems Answers are included in the Appendix Science Books & Films, 1977

Chemistry John S. Phillips, Cheryl Wistrom, 2000 Prentice Hall Physical Science Concepts in Action Program Planner National Chemistry Physics Earth Science, 2003-11 Prentice Hall Physical Science Concepts in Action helps students make the important connection between the science they read and what they experience every day Relevant content lively explorations and a wealth of hands on activities take students understanding of science beyond the page and into the world around them Now includes even more technology tools and activities to support differentiated instruction **The Software Encyclopedia 2000** Bowker Editorial Staff, 2000-05 **Computational Methods in Science and Engineering** George Maroulis, Theodore E. Simos, 2007-12-14 All papers have been peer reviewed The aim of ICCMSE 2007 is to bring together computational scientists and engineers from several disciplines in order to share methods methodologies and ideas The potential readers of these proceedings are all the scientists with interest in the following fields Computational Mathematics Theoretical Physics Computational Physics Theoretical Chemistry Computational Chemistry Mathematical Chemistry Computational Engineering Computational Mechanics Computational Biology and Medicine Scientific Computation High Performance Computing Parallel and Distributed Computing Visualization Problem Solving Environments Software Tools Advanced Numerical Algorithms Modeling and Simulation of Complex Systems Web based Simulation and Computing Grid based Simulation and Computing Computational Grids and Computer Science Holt Chemistry Ralph Thomas Myers, 2004

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