

Lab Report:
Analysis of an aluminum-zinc alloy

CHEM 200 section:1

Analysis of an aluminum-zinc alloy

Experimental Procedures

Nothing changed from pre-lab procedures

Part A Data:

Room temperature (Celcius)	21
Mass of aluminum (g)	0.030
Initial buret reading (mL)	21.6
Final Buret Reading (mL)	56.3
Height of water column (cm)	29.2

Calculations:

Volume of H₂ (mL)

$$56.3\text{mL} - 21.6\text{mL} = 34.7\text{mL}/1000 \\ = 0.0347 \text{ L}$$

Pressure of H₂O from room temperature (atm)=0.0231 atm

Partial pressure of H₂ (atm): $P_{H_2} = p_{\text{atm}} - 0.987 \left(\frac{1 \text{ atm}}{1026 \text{ cm}} \times 29.2 \right) - 0.0231 = 0.512 \text{ atm}$

H₂ gas actual yield:

R=0.0821atm

T=(21+273.15)=294.15

$$n_{H_2} = \frac{(0.512 \text{ atm})(0.0347 \text{ L})}{(0.0821 \text{ atm})(294.15 \text{ K})} = 0.00736 \text{ mol}$$

Balanced equation:



Theoretical yield:

$$\frac{(\text{mass of Al})(1 \text{ mol Al})(3 \text{ mol of H}_2)}{(\text{molar mass of Al})(2 \text{ mol Al})} = \text{theoretical yield} \\ \frac{(0.030 \text{ g Al})(1 \text{ mol Al})(3 \text{ mol Al})}{(26.98 \text{ g/mol Al})(2 \text{ mol Al})} = 0.00278 \text{ mol H}_2(\text{g})$$

Percent Yield:

$$\frac{\text{actual yield}}{\text{theoretical yield}} \times 100 = \frac{0.00736 \text{ mol}}{0.00278 \text{ mol}} \times 100 = 64.7\%$$

Part B Data:

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