

CHEMISTRY FORM 4 SCHEMES OF WORK TERM 1,2,3

FORM FOUR CHEMISTRY							
SCHEME OF WORK				TERM ONE			
WK NO	LJ NO	TOPIC / SUBTOPIC	LESSON / SPECIFIC OBJECTIVES	TEACHING / LEARNING ACTIVITIES	MATERIALS / RESOURCES	REFERENCES	REMARKS
1		ACIDS, BASES AND SALTS.	<u><i>By the end of the lesson, the learner should be able to</i></u>				
	1.	Strength of acids.	Define an acid in terms of hydrogen ions.	Class experiments: investigate reactions of magnesium and zinc carbonate with different acids. Make and record observations in tabular form.	Magnesium strip, zinc carbonate, 2M HCl, 2M H ₂ SO ₄ , 2M ethanoic acid.	<i>K.L.B. BK IV Pages 1-4</i>	
	2	Acids in aqueous form.	Explain strength of acids in aqueous form in terms of number of hydrogen ions present.	Make deductions from the observations. Write relevant chemical equations and ionic equations. Detailed discussion leading to the definition of an acid and explanation of strength of an acid.			
	3.	pH values of acids.	Determine strength of acids using pH values.	Q/A: review determination of strength of acids using a litmus paper and pH scale. Class / group experiments: record colour of universal indicator in 2M HCl and 2M ethanoic acid.	Universal indicator, 2M HCl, 2M ethanoic acid, dry cells, carbon electrodes, milli-ammmeters, wires, switches etc.	<i>K.L.B. BK IV Pages 4-6</i>	
	4	Electrical conductivities of aqueous acids.	Determine strengths of acids by comparing their electrical conductivities. Classify acids as either strong or weak in terms of partial dissociations in aqueous solutions.	Set up voltmeters of 2M HCl and 2M ethanoic acid in turn. Record ammeter of current . Discuss the observations. Write corresponding ionic equations.			
	5	Definition of a base in terms of hydroxide ions.	Define a base in terms of hydroxide ions.	Teacher demonstration: Dissolve calcium hydroxide in water. Carry out litmus test on the resulting solution. Discuss the results; hence define a base in terms of hydroxide ions.	Red litmus paper, calcium hydroxide solid.	<i>K.L.B. BK IV Pages 6-7</i>	

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Scheme Of Work Of Chemistry For Third Term2015

Tobias Bleicker



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Table of Contents Scheme Of Work Of Chemistry For Third Term2015

1. Understanding the eBook Scheme Of Work Of Chemistry For Third Term2015
 - The Rise of Digital Reading Scheme Of Work Of Chemistry For Third Term2015
 - Advantages of eBooks Over Traditional Books
2. Identifying Scheme Of Work Of Chemistry For Third Term2015
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Scheme Of Work Of Chemistry For Third Term2015
 - User-Friendly Interface
4. Exploring eBook Recommendations from Scheme Of Work Of Chemistry For Third Term2015
 - Personalized Recommendations
 - Scheme Of Work Of Chemistry For Third Term2015 User Reviews and Ratings
 - Scheme Of Work Of Chemistry For Third Term2015 and Bestseller Lists
5. Accessing Scheme Of Work Of Chemistry For Third Term2015 Free and Paid eBooks
 - Scheme Of Work Of Chemistry For Third Term2015 Public Domain eBooks
 - Scheme Of Work Of Chemistry For Third Term2015 eBook Subscription Services
 - Scheme Of Work Of Chemistry For Third Term2015 Budget-Friendly Options

6. Navigating Scheme Of Work Of Chemistry For Third Term2015 eBook Formats
 - ePub, PDF, MOBI, and More
 - Scheme Of Work Of Chemistry For Third Term2015 Compatibility with Devices
 - Scheme Of Work Of Chemistry For Third Term2015 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Scheme Of Work Of Chemistry For Third Term2015
 - Highlighting and Note-Taking Scheme Of Work Of Chemistry For Third Term2015
 - Interactive Elements Scheme Of Work Of Chemistry For Third Term2015
8. Staying Engaged with Scheme Of Work Of Chemistry For Third Term2015
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Scheme Of Work Of Chemistry For Third Term2015
9. Balancing eBooks and Physical Books Scheme Of Work Of Chemistry For Third Term2015
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Scheme Of Work Of Chemistry For Third Term2015
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Scheme Of Work Of Chemistry For Third Term2015
 - Setting Reading Goals Scheme Of Work Of Chemistry For Third Term2015
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Scheme Of Work Of Chemistry For Third Term2015
 - Fact-Checking eBook Content of Scheme Of Work Of Chemistry For Third Term2015
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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