Connecting to Your World

You know that without oxygen. to breathe, you could not live. But did you know that on gen plays

another important role in your life? High in the atmosphere, a different form of payeen, called ozone, forms a layer that filters.

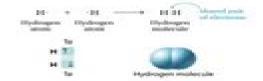
> out harmful radiation from the sun. The colors in this map indicate the concentrations of ozone in various parts of Earth's atmosphere in this section, you will learn how oxygen atoms canjoin in pairs to form the oxygen you breathe and can also join in groups of three oxygen atoms to florein gazone.

The Octet Rule in Covalent Bonding

Recall that when ionic compounds form, electrons tend to be transferred so that each ion acquires a noble gas configuration. A similar rule applies for consient bonds. 😂 in covalent bonds, electron sharing usually occurs so that atoms attain the electron configurations of poblic gases. For example, each hydrogen atom has one electron. But a pair of hydrogen atoms share these two electrons when they form a covalent bond in a hydrogen molecule. Each hydrogen atom thus attains the electron configuration of belium, a noble gas with two electrons. Combinations of atoms of the nonmetals and metalloids in Groups 6A, 5A, 6A, and 7A of the periodic table are likely to form covalent bonds. In this case the atoms usually acquire a total of eight electrons, or an octet, by sharing electrons, so that the occet rule applies.

Single Covalent Bonds

The hydrogen anoms in a hydrogen molecule are held together mainly by the attraction of the shared electrons to the positive model. Two atoms held together by sharing a pair of electrons are joined by a single savalent bond. Hydrogen gas consists of distornic molecules whose atoms share only one pair of electrons, forming a single covalent bond.



Guide for Reading

Key Concepts

- What is the result of electron. Tabolod Involuncy in goleans
- How do electron dot structures
- represent shared electrons? How do attorns form double or
- telephonological beyonds/1
- How are considerate considerat Name And Address of Pages and Address Telephonel Impeliance
- 9. More in the strength of a allowalterni blowed included to its Typewal allowardation emergy?
- How are oxygen atoms borded. the management
- 'Bifruit are some escaptions to titles account make?"

Vocabulary

sangle covalent bond stimuctural formula weeksmed pair About the Considerate Balance tengths deplusioned bound. coordinate covalent bond polystomic sec bond dissociation energy personal property of the party of the party

Reading Strategy

Identifying Main Idea/Details Cost title engine billion to their marks. bright following the heading The Octot Rule in Counters Bonding As your read, list expension of Roose Office make its Softlineard suffrence strigle covalent band, a double consident blood, and a triple cova-Servet Second House.

Section 8.2. The Nature of Covalent Bonding

Reading For 8 2 The Nature Of Covalent Bonding

Judith Kinnear, Marjory Martin, Lucy Cassar, Elise Meehan, Ritu Tyagi

Reading For 8 2 The Nature Of Covalent Bonding:

Symmetry and Physical Properties of Crystals Cécile Malgrange, Christian Ricolleau, Michel Schlenker, 2014-12-04
Crystals are everywhere from natural crystals minerals through the semiconductors and magnetic materials in electronic devices and computers or piezoelectric resonators at the heart of our quartz watches to electro optical devices
Understanding them in depth is essential both for pure research and for their applications. This book provides a clear thorough presentation of their symmetry both at the microscopic space group level and the macroscopic point group level. The implications of the symmetry of crystals for their physical properties are then presented together with their mathematical description in terms of tensors. The conditions on the symmetry of a crystal for a given property to exist then become clear as does the symmetry of the property. The geometrical representation of tensor quantities or properties is presented and its use in determining important relationships emphasized. An original feature of this book is that most chapters include exercises with complete solutions. This allows readers to test and improve their understanding of the material. The intended readership includes undergraduate and graduate students in materials science and materials related aspects of electrical and optical engineering researchers involved in the investigation of the physical properties of crystals and the design of applications based on crystal properties such as piezoelectricity electro optics optical activity and all those involved in the characterization of the structural properties of materials.

University of Michigan Official Publication, 1949

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