

1. True or false? (5 marks)

- An atom has the same number of protons as it does electrons. *F (true if the atom is neutral, false because an atom doesn't have to be neutral)*
- The atomic number is the number of protons in an atom. *T*
- An atom always has the same number of neutrons as it does protons. *F*
- The mass number is the number of neutrons in an atom. *F*
- The mass number is the number of particles in the nucleus. *T*

2. Complete the following table (5 marks)

Subatomic particle	Letter used as symbol	Charge	Mass	Location
Proton	<i>p</i>	<i>+1</i>	<i>1000 times heavier than an electron</i>	<i>Nucleus</i>
Neutron	<i>n</i>	<i>none</i>	<i>1000 times heavier than an electron</i>	<i>Nucleus</i>
electron	<i>e</i>	<i>-1</i>	<i>Almost negligible</i>	<i>Electron orbital/shells</i>

3. Draw a picture of an Oxygen atom showing the: (5 marks)

- 8 protons
- 8 neutrons
- 8 electrons
- The nucleus
- Electron shells



4. Pictured are all the elements you need to create a timeline for the model of atomic structure. Use the blank timeline to write years and beside each year write the important discovery at that time. (4 marks)



5. Four groups all measured the boiling point of water. Their measurements were: (4 marks)

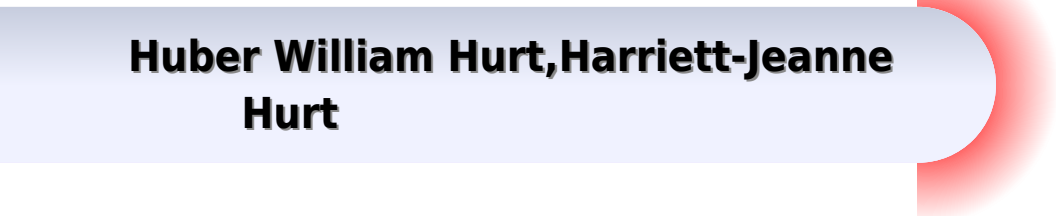
Anth's group	100.1°C
Fiona's group	98.9°C
Joe's group	75.2°C
Shamika's group	101.3°C

After the experiment, Joe claimed that the boiling point of water was approximately 76°C.

- Were his experimental observations replicated? *no*
- What do you make of his claim? *It is not supported by the rest of the experimental evidence, so it would reject his hypothesis.*
- According to this data, calculate the boiling point of water $\frac{100.1 + 98.9 + 101.3}{3} = 100.1^\circ\text{C}$ *or 100.0°C if Joe's data is included.*

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**Huber William Hurt, Harriett-Jeanne
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