

Science

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(Chapter – 4) (Structure of the Atom) (Exercises)

(Class – IX)

Question 1:

Compare the properties of electrons, protons and neutrons.

Answer 1:

Electron	Proton	Neutron
(i) Electron are present outside the nucleus of an atom.	(i) Proton are present in the nucleus of an atom.	(i) Neutron are present in the nucleus of an atom.
(ii) Electron are negatively charged	(ii) Proton are positively charged.	(ii) Neutron are neutral.
(iii) The mass of an electron is considered to negligible.	(iii) The mass of a proton is approximately 2000 times as the mass of an electron	(iii) The mass of neutron is nearly equal to the mass of a proton.

Question 2:

What are the limitations of J.J. Thomson's model of the atom?

Answer 2:

According to J.J. Thomson's model of an atom, an atom consists of a positively charged sphere with electrons embedded in it. However, it was later found that the positively charged particles reside at the center of the atom called the nucleus, and the electrons revolve around the nucleus.

Question 3:

What are the limitations of Rutherford's model of the atom?

Answer 3:

According to Rutherford's model of an atom, electrons revolve around the nucleus in fixed orbits. But, an electron revolving in circular orbits will not be stable because during revolution, it will experience acceleration. Due to acceleration, the electrons will lose energy in the form of radiation and fall into the nucleus. In such a case, the atom would be highly unstable and collapse.

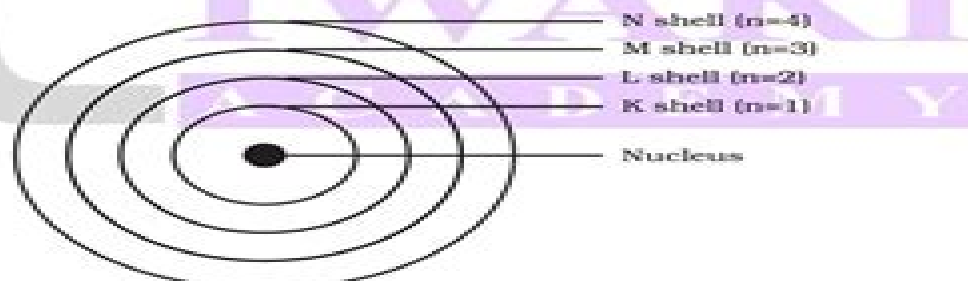
Question 4:

Describe Bohr's model of the atom.

Answer 4:

Bohr's model of the atom Niels Bohr proposed the following postulates regarding the model of the atom.

- Only certain orbits known as discrete orbits of electrons are allowed inside the atom.
- While revolving in these discrete orbits, the electrons do not radiate energy. These discrete orbits or shells are shown in the following diagram.



The first orbit (i.e., for $n = 1$) is represented by letter K. Similarly, for $n = 2$, it is L - shell, for $n = 3$, it is M - shell and for $n = 4$, it is N - shell. These orbits or shells are also called energy levels.

Science 9 Chapter 4 Test

Lisa Mc Coll, Leigh Angela Brady

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located on the right side of the slide.

Science 9 Chapter 4 Test:

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Experimental Economics Nicholas Bardsley,Robin Cubitt,Graham Loomes,Peter Moffatt,Chris Starmer,Robert Sugden,2020-05-26 Since the 1980s there has been explosive growth in the use of experimental methods in economics leading to exciting developments in economic theory and policy Despite this the status of experimental economics remains controversial In *Experimental Economics* the authors draw on their experience and expertise in experimental economics economic theory the methodology of economics philosophy of science and the econometrics of experimental data to offer a balanced and integrated look at the nature and reliability of claims based on experimental research The authors explore the history of experiments in economics provide examples of different types of experiments and show that the growing use of experimental methods is transforming economics into a genuinely empirical science They explain that progress is being held back by an uncritical acceptance of folk wisdom regarding how experiments should be conducted a failure to acknowledge that different objectives call for different approaches to experimental design and a misplaced assumption that principles of good practice in theoretical modeling can be transferred directly to experimental design *Experimental Economics* debates how such limitations might be overcome and will interest practicing experimental economists nonexperimental economists wanting to interpret experimental research and philosophers of science concerned with the status of knowledge claims in economics

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