

2017-2018

School Employee Appreciation Dates

September 19, 2017	National IT Professional Day
October 2, 2017	National Custodial Appreciation Day
October 9-13, 2017	National School Lunch Week
November 17, 2017	Substitute Educator Day
November 13-17, 2017	American Education Week
November 19, 2017	National Educational Support Professional Day*
December 4-8, 2017	Maintenance Employee Appreciation Week
January 17, 2018	National Mentoring Day
February 6-9, 2018	National School Counselor Week
March 5-11, 2018	School Social Worker Week
April 17-21, 2018	Public School Volunteer Week
April 23-29, 2018	Administrative Professional Week
April 4, 2018	Media Specialist Day
April 25, 2018	Administrative Professional Day
April 30-May 4, 2018	National Child Nutrition Employee Week
May 1, 2018	School Principals Day
May 7-11, 2018	Orndorff County Schools Employee Appreciation Week
May 7-11, 2018	Orndorff County Schools Bus Driver Appreciation Week
May 7-11, 2018	Teacher & Teacher Assistant Appreciation Week
May 8, 2018	National Teacher Day
May 9, 2018	National School Nurse Day
May 13, 2018	Teacher Assistant Appreciation Day*
May 20-26, 2018	National Educational Boss Week

*Includes Speech Therapists, School Psychologists, Program Specialists, Occupational Therapists and Assistants, Physical Therapists and Assistants, Educational Interpreters, and any other employees not recognized elsewhere.

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web with a strong acid and strong base titration strong acids and strong bases refer to species that completely dissociate to form ions in solution by contrast weak acids and bases ionize only partially and the ionization reaction is reversible

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web conclusion the purpose of the experiment was to identify the unknown solid acid 17 by calculating its molar mass this was done by titration of the acid with sodium hydroxide in order to find the volume of sodium hydroxide required to reach the end point and that value was used in unit analysis to find the molar mass and identify the unknown solid acid

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web by utilizing the half equivalence point on the titration curve the pka of the acid analyte was determined to be 3.655
furthermore using the stoichiometric ratio between the naoh and the acid analyte 1:1 we could experimentally determine the number of moles of the acid in the solution to be 0.00731 moles making the experimental molar mass

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 $HCl(aq) + NaOH(aq) \rightarrow NaCl(aq) + H_2O(l)$
one mole of HCl would be fully neutralized by one mole of NaOH

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a pH titration is used to determine the concentration of an acid or a base titrations play an important role in determining amount and purity in many manufacturing processes

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web specifically an acid base titration can be used to figure out the concentration of an acid or base whether an unknown acid or base is strong or weak pK_a of an unknown acid or pK_b of the unknown base key terms titration a process where a solution of known strength is added to a certain volume of a treated sample containing an

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web conclusion this lab evaluated an acid base reaction with the volumetric analysis of a standard solution being naoh reaction with a 2nd substance $\text{H}_2\text{C}_2\text{O}_4$ aq 2NaOH aq $\text{Na}_2\text{C}_2\text{O}_4$ $4\text{H}_2\text{O}$ to differentiate an acid from a base in a solution a ph scale is used the ph scale ranges from 0 to 14 a ph of 7 is neutral in an acid base reaction

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web this point of titration is referred to as the equivalence point considering the 1 1 stoichiometry of this acid base reaction NaOH aq $\text{C}_6\text{H}_4\text{COOH}$ aq $\text{C}_6\text{H}_4\text{COONa}$ aq H_2O l the point of equivalence is the point of titration when the number of moles of naoh na added is equal to the number of moles of khp nb in the

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