

SCH4C Practice WS Unit 1**Multiple Choice***Identify the choice that best completes the statement or answers the question.*

- _____ 1. The special band of light waves that the human eye can detect is known as the
a. continuous spectrum c. radio spectrum
b. line spectrum d. visible spectrum
- _____ 2. Electromagnetic energy is
a. a measure of a magnetic field c. electric energy
b. light energy d. heat energy
- _____ 3. Which of the following is an example of light energy?
a. visible light c. ultraviolet light
b. X-rays d. all of the above
- _____ 4. Each element produces a characteristic
a. line spectrum c. visible spectrum
b. continuous spectrum d. none of the above
- _____ 5. For light that is emitted in the visible range, the colour of the light depends upon
a. the type of spectroscope used c. the element that is used
b. the frequency of the light wave d. all of the above
- _____ 6. According to the law of electric charges, Rutherford's "atom" should have
a. emitted ultraviolet light c. collapsed
b. been very stable d. emitted X-rays
- _____ 7. Niels Bohr was able to explain the line spectrum of
a. helium c. neon
b. hydrogen d. sodium
- _____ 8. Bohr's theory is based on the fact that an electron's energy is
a. negligible c. always the same
b. very large d. quantized
- _____ 9. Bohr compared the electrons in an atom to
a. raisins in a bun c. planets orbiting the sun
b. bees around a beehive d. stars in the sky
- _____ 10. When an electron gains a specific amount of energy,
a. it moves to a higher energy level c. it moves to its ground state
b. it gives off light energy d. none of the above
- _____ 11. An electron that occupies a higher energy level than normal is said to be in
a. the wrong energy level c. a higher state
b. an excited state d. the ground state
- _____ 12. Electrons in their ground state
a. are not moving c. do not emit any light energy
b. give off light energy continually d. none of the above

Sch4c Environmental Chemistry Test

Gerard W. Vollmer



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