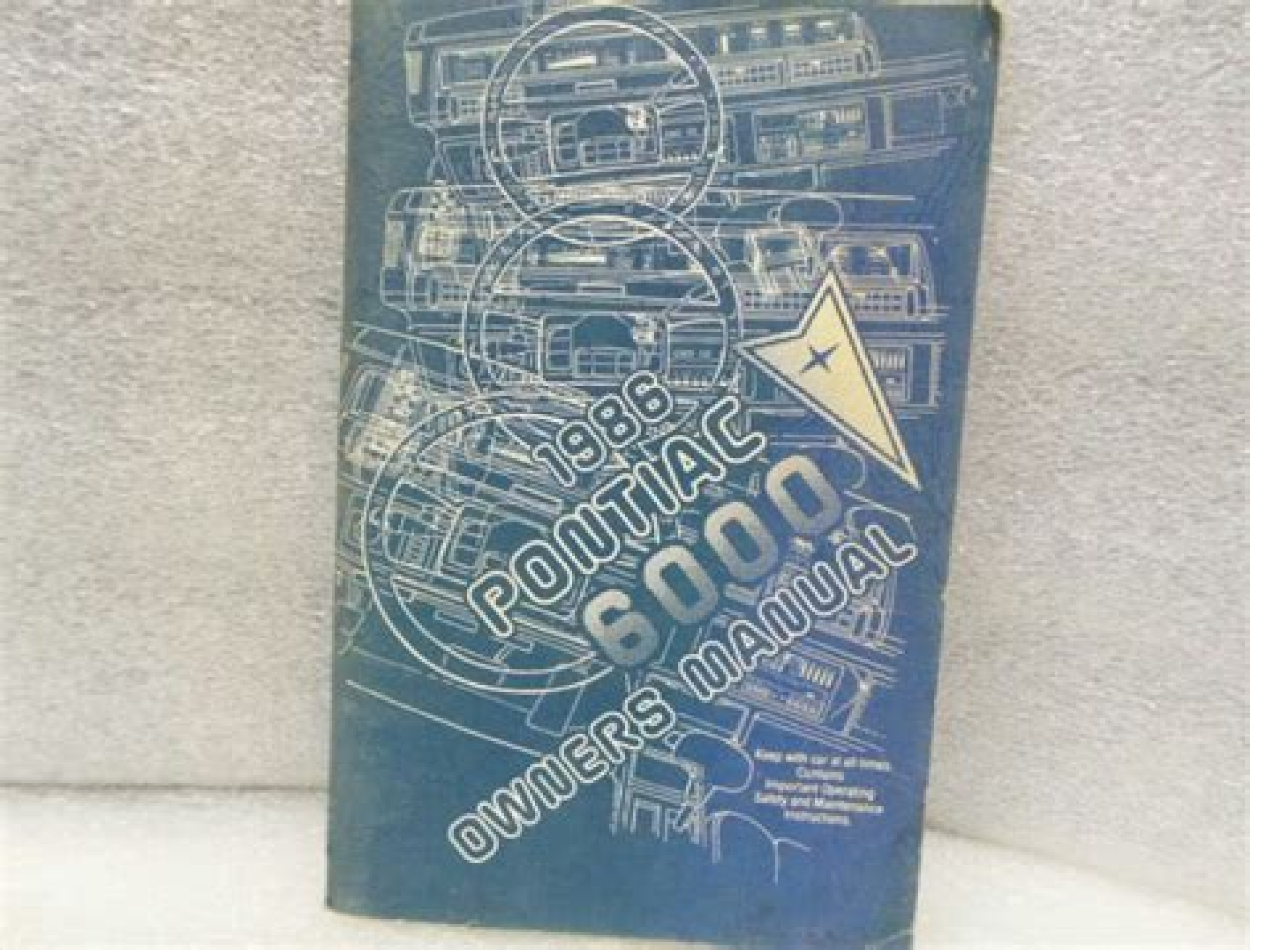


The background of the manual cover is a dark blue technical illustration of a Pontiac 6000. The car is shown from a front-three-quarter perspective. Several circular callouts are overlaid on the car, highlighting specific components like the engine, transmission, and suspension. A yellow star with a black crosshair is positioned to the right of the car's front end.

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web consider the thermochemical equation below

$$4\text{NO(g)} + 6\text{H}_2\text{O(l)} + 4\text{NH}_3\text{(g)} + 5\text{O}_2\text{(g)} \rightarrow 11\text{H}_2\text{O(l)} + 7\text{N}_2\text{(g)}$$
 is this reaction endothermic or exothermic
 endothermic
 how much heat will be absorbed or released if 10 g of NO is reacted with 10 g of H₂O

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web heats of fusion and vaporization determines what determines how much energy is required to change a substance's phase
 what is the heat of fusion formula
 $q = m \cdot \Delta H_{\text{fus}}$
 what is the heat of vaporization formula
 $q = m \cdot \Delta H_{\text{vap}}$
 what is on the x-axis of the heating curve
 time

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 3 carry out a thermochemical calculation on all these species at their optimized geometries
 4 sum up the enthalpies free energy of the products and the reagents keeping into account the

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web answer a
 1 explanation the entire reaction was multiplied by 2 therefore ΔH must also be halved
 3 answer c
 explanation option c is incorrect because 566 kJ/mol is the ΔH per mole of O₂ and not for 2 moles even though options b and d are expressed in grams they are still correct because 56

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 what is the heat change when 472 g of carbon reacts with excess O₂ according to the following equation
 how much heat energy must be expelled by the body to rid the body of one pound of fat
 2530 J
 2530 J
 3 h 7914 kJ is this endothermic or exothermic
 2

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 exothermic transfer of heat from the system to the surroundings
 2 h 2 g O₂ g 2 h 2 O l energy endothermic the transfer of heat from the surroundings to the system
 energy 2 h g s 2 h g l O₂ g

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 enthalpy is an extensive property depends on amount of matter present this fact means that the magnitude of h is directly proportional to the amount of reactant consumed in the process

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web q 140 g 4 184 J 160 C g O₂ q 9372 J 16 J 35 5 g x 1 mole KCN 0 5453 moles 65 1 g 0 5453 mole moles 5
 when a 15 g sample of SrCl₂ dissolves in 80 g of water in a calorimeter the temperature drops from 42 C to 15 C calculate the ΔH for the process

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