

REPORT FOR EXPERIMENT 12

INSTRUCTOR

Single Displacement Reactions

Evidence of Reaction	Equations (to be completed)
Describe any evidence of reaction; if no reaction was observed, write "None".	Write "No reaction", if no reaction was observed.
1. <u>spongy, dark red/black</u>	$\text{Cu} + \text{AgNO}_3(\text{aq}) \rightarrow \text{Cu(NO}_3)_2 + 2\text{Ag}$
2. <u>dark</u>	$\text{Pb} + \text{Cu(NO}_3)_2(\text{aq}) \rightarrow \text{Pb(NO}_3)_2 + \text{Cu}$
3. <u>spongy, dark</u>	$\text{Zn} + \text{Pb(NO}_3)_2(\text{aq}) \rightarrow \text{Zn(NO}_3)_2 + \text{Pb}$
4. <u>none</u>	$\text{Zn} + \text{MgSO}_4(\text{aq}) \rightarrow \text{no reaction}$
5. <u>none</u>	$\text{Cu} + \text{H}_2\text{SO}_4(\text{aq}) \rightarrow \text{no reaction}$
6. <u>same color spongy being formed</u>	$\text{Zn} + \text{H}_2\text{SO}_4(\text{aq}) \rightarrow \text{ZnSO}_4 + \text{H}_2$

QUESTIONS AND PROBLEMS

- Complete the following table by writing the symbols of the two elements whose reactivities are being compared in each test:

	Tube Number					
	1	2	3	4	5	6
Greater activity	<u>Cu</u>	<u>Pb</u>	<u>Zn</u>	<u>Mg</u>	<u>H</u>	<u>Zn</u>
Lesser activity	<u>Ag</u>	<u>Cu</u>	<u>Pb</u>	<u>Zn</u>	<u>Cu</u>	<u>H</u>

- Arrange Pb, Mg, and Zn in order of their activities, listing the most active first.

- Mg
- Zn
- Pb

Report For Experiment 12 Single Displacement Reactions

Terry C. Jones



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