

Answer following



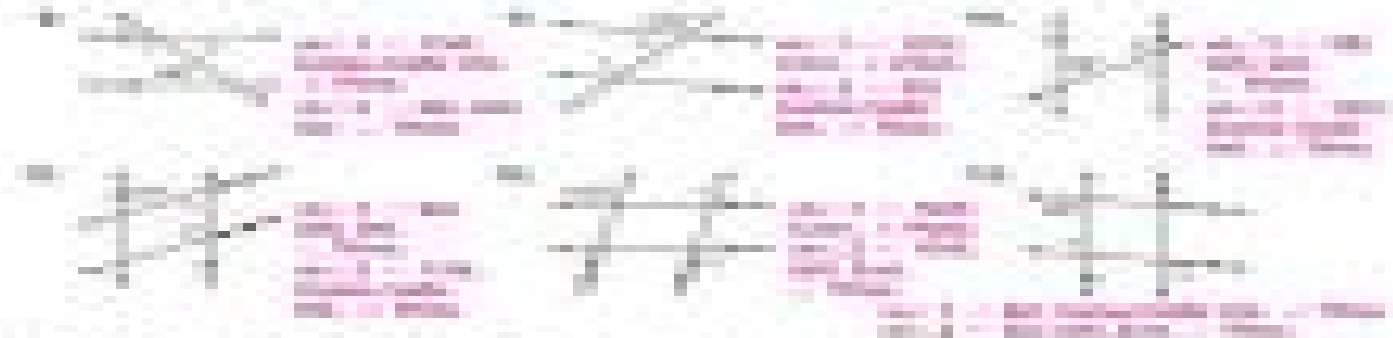
- 1. Substrate and temperature: 20°C $100\text{ mg} \cdot \text{min}^{-1} \cdot \text{mg}^{-1}$
- 2. A graph of product vs. time: $100\text{ mg} \cdot \text{min}^{-1} \cdot \text{mg}^{-1}$ $100\text{ mg} \cdot \text{min}^{-1} \cdot \text{mg}^{-1}$
- 3. A graph of substrate vs. time: $100\text{ mg} \cdot \text{min}^{-1} \cdot \text{mg}^{-1}$ $100\text{ mg} \cdot \text{min}^{-1} \cdot \text{mg}^{-1}$

Calculate all values of K_m and V_{max} for the following reaction. Assume that the reaction is reversible and that the reaction is at equilibrium.

- 1. $K_m = 100\text{ mg} \cdot \text{min}^{-1} \cdot \text{mg}^{-1}$ $V_{max} = 100\text{ mg} \cdot \text{min}^{-1} \cdot \text{mg}^{-1}$
- 2. $K_m = 100\text{ mg} \cdot \text{min}^{-1} \cdot \text{mg}^{-1}$ $V_{max} = 100\text{ mg} \cdot \text{min}^{-1} \cdot \text{mg}^{-1}$
- 3. $K_m = 100\text{ mg} \cdot \text{min}^{-1} \cdot \text{mg}^{-1}$ $V_{max} = 100\text{ mg} \cdot \text{min}^{-1} \cdot \text{mg}^{-1}$
- 4. $K_m = 100\text{ mg} \cdot \text{min}^{-1} \cdot \text{mg}^{-1}$ $V_{max} = 100\text{ mg} \cdot \text{min}^{-1} \cdot \text{mg}^{-1}$

Answer: 100

Graph the following data. Assume the reaction is reversible and that the reaction is at equilibrium.



Graph the following data. Assume the reaction is reversible and that the reaction is at equilibrium.



Graph the following data. Assume the reaction is reversible and that the reaction is at equilibrium.

Prentice Hall Geometry 4 5 Answers

Liying Dong



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