Answers for Predicting Products of Chemical Reactions

For all combustion reactions of hydrocarbons or hydrogen-carbon-oxygen molecules, the products will always be CO: and H:O (assuming a complete reaction).

a. Combustion:	C_6H_{12}	+	90a	-	6CO ₂		6H;O
b. Combustion:	2C4He	+	1102	\rightarrow	8CO2	+	6H ₂ O
c. Combustion:	$C_6H_{10}O_3$	94	707		6CO ₂		5H ₂ O

1. Synthesis: Mg + I₂ → MgI₁

Note that Mg can only form Mg2 and I can only form I', so the product will must have a 1:2 cation anion ratio.

Double displacement: CuCl₂ + H₂S → CuS + 2HCl
 Note that the product is not H₂Cl₂. It is important to recognize that CuCl₂ is made of three ions, Cu²⁺ and two Cl².

Double displacement: NaOH + HCIO₄ → NaCIO₄ + H₂O
 In this question, you must recognize that perchlorate, ClO₄, and hydroxide, OH, are polyatomic ions and will not break apart. Also, this is an acid-base reaction, so the products should be salt and water.

Decomposition: ZnCO₂ + heat → ZnO + CO₂.
 When reactions have heat as a reactant, it is very likely that they will involve decompositions. Carbonate compounds usually decompose to CO₂ and a metal oxide.

5. Single replacement: 2HCl + Zn → ZnCl₂ + H₂
Note that one reactant is in its elemental form; if a single replacement reaction is going to occur, the species NOT in its elemental form in the reactants (H⁺ in this case), will end up in its elemental form in the products (H₂) and the species that IS in its elemental form (Zn) will end up ionized (Zn²⁺). Note that zinc can only form a Zn²⁺ ion, so it will have two chloride ions. Note also that hydrogen in its elemental form is H₂, not H.

6. Single replacement 2Na + MgCl₂ → 2NaCl + Mg Again notice that one species is in its elemental form (Na). The magnesium in MgCl₂ is an ion (Mg²¹), but is transformed into its elemental state (Mg), while the Na is converted into an ion (Na¹, sodium only forms a +1 ion).

Double replacement CsCl: + K:CO → CsCO: + 2KCl
 Recognize that carbonate is a polyatomic ion (CO₁²) and that the cations are already stable ions (Ca²⁺ and K²).

8. <u>Synthesis</u> 2K + Cl₂ → 2KCl

Note that both materials are elemental species, so the only result could be a synthesis reaction. In the product, K^{*} and CT are formed. Note that we would NOT write the product as K₂Cl₂. Ionic compounds are written in their simplest forms.

Double replacement 3BaCl: + 2K:PO₁ → Bar(PO₁): + 6KCl
 Note that phosphate (PO₁) is a polyatomic ion and will not break apart. Since barium is a +2 ion, the barium phosphate will have a 3:2 ratio of BarPO₁ in order to balance the charge.

10. <u>Double replacement</u> H:SO_a + 2KOH → 2H:O + K:SO_a Notice that sulfate (SO_a²) is a polyatomic ion and that potassium exists as a +1 ion (K*). Also, this is an acid-base reaction, so the products should be salt and water.

Decomposition
 Al:(CO₂): + beat → Al:O₁ + 3CO₂
 Another decomposition that will generate CO₂ and a metal oxide (note that Al forms a +3 ion, Al²⁺, and monatomic oxygen will have a −2 charge, O²⁺).

Predicting Products Of Chemical Reactions Answers

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