

16. The dissociation of water is shown below:



- At 25 °C, the auto product of water, K_w , is 1.0×10^{-14} mol² dm⁻⁶.
At 65 °C, the auto product of water, K_w , is 1.0×10^{-13} mol² dm⁻⁶.

17. Explain what the dissociation of water is an indication of water's behavior.

Endothermic.

18. Predict, using a calculation, whether a solution of 1.0 mol dm⁻³ is acidic, basic or neutral. (3)

Acidic

Explain by calculating pH.

19. pK_a , pH and pOH are significant notions.

Calculate pK_a of H_2O .

Give your answer to two decimal places.

$$-\log(1.0 \times 10^{-14}) = 14.00$$

$$pK_a = 14.00$$

20. Describe the autoionization of water in terms of water and the resulting species in solution at 25 °C.

Calculate the pH of the resulting solution of water at 25 °C.

Give your answer to two decimal places.

$$pH = 7.00$$

pH = _____ (3)

(Total: 18)

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Richard Parsons



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